

1 **Gliding arc plasma for CO₂ conversion: better insights by a combined** 2 **experimental and modelling approach**

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9 **Highlights**

- 10 ✚ A two dimensional self-consistent model is developed and validated by the direct experiment.
- 11 ✚ Gliding arc shows a strong non-equilibrium character of the conversion process, explaining the
12 higher values of conversion and energy efficiency than thermal process.
- 13 ✚ A chemical kinetics analysis shows that the CO₂ vibrational levels significantly contribute to the
14 CO₂ dissociation.
- 15 ✚ Promoting the vibrational kinetics, reducing the recombination of CO with O₂ and increasing the
16 CO₂ fraction treated by the arc can further improve the conversion and energy efficiency.

17 **Abstract**

18 A gliding arc plasma is a potential way to convert CO₂ into CO and O₂, due to its non-equilibrium
19 character, but little is known about the underlying mechanisms. In this paper, a self-consistent two-
20 dimensional (2D) gliding arc model is developed, with a detailed non-equilibrium CO₂ plasma
21 chemistry, and validated with experiments. Our calculated values of the electron number density in
22 the plasma, the CO₂ conversion and energy efficiency show reasonable agreement with the
23 experiments, indicating that the model can provide a realistic picture of the plasma chemistry.
24 Comparison of the results with classical thermal conversion, as well as other plasma-based
25 technologies for CO₂ conversion reported in literature, demonstrates the non-equilibrium character
26 of the gliding arc, and indicates that the gliding arc is a promising plasma reactor for CO₂ conversion.
27 However, some process modifications should be exploited to further improve its performance. As the
28 model provides a realistic picture of the plasma behaviour, we use it first to investigate the plasma
29 characteristics in a whole gliding arc cycle, which is necessary to understand the underlying
30 mechanisms. Subsequently, we perform a chemical kinetics analysis, to investigate the different
31 pathways for CO₂ loss and formation. Based on the revealed discharge properties and the underlying
32 CO₂ plasma chemistry, the model allows us to propose solutions on how to further improve the CO₂
33 conversion and energy efficiency by a gliding arc plasma.

34 **Keywords:** CO₂ conversion, gliding arc, non-equilibrium plasma, plasma chemistry, splitting
35 mechanisms, breakdown

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1 **1. Introduction**

2 Plasma technology offers unique perspectives, because of its capacity to induce chemical
3 reactions within gases at ambient temperature and pressure, due to its non-equilibrium character.
4 Plasma is created by applying electric power to a gas, causing breakdown of the gas into ions and
5 electrons and also producing a large number of reactive species, such as various radicals and excited
6 species. This makes plasma a highly reactive cocktail, which is quite promising for greenhouse gas
7 conversion. Indeed, the inert CO₂ gas is activated by electron impact ionization, excitation and
8 dissociation. Furthermore, plasma is very flexible and can easily be switched on and off, so it is quite
9 promising for storing peak renewable energy into fuels. Indeed, more and more electrical energy
10 nowadays is produced from renewable energy sources (wind or solar), which often suffer from
11 fluctuating peak powers, making it difficult to match the supply of this electricity with the demand.
12 This surplus of electricity can in principle be used in plasma to convert greenhouse gases into value
13 added chemicals when adding a suitable H-source to the CO₂ gas, such as H₂O, CH₄ or H₂. However,
14 there is still a long way to go, certainly if we target the selective production of some value-added
15 products, for which the combination with a suitable catalyst would be needed. This makes plasma
16 based greenhouse gas conversion fit in principle in the framework of green chemistry [1]-[2] and also
17 complies with the “cradle-to-cradle” principle [3].

18 Gliding arc (GA) plasmas are potential plasma sources for gas conversion [4]-[18] because they
19 offer benefits of both thermal and non-thermal discharges. They are typically considered as ‘warm’
20 discharges, which are characterized by a better energy efficiency than other types of plasmas. The
21 reason is that they provide efficient vibrational excitation of the molecules, which is seen as the most
22 energy-efficient way to split CO₂ molecules [19].

23 In order to improve the applications (i.e., mainly gas conversion), the physical and chemical
24 characteristics of the GA have been extensively studied by experiments, including high-speed
25 photography [20], electrical measurements [21]-[23] and spectroscopic measurements [24]-[25].
26 Besides experiments, detailed modelling is also very useful to provide more insight into the
27 underlying reaction mechanisms of plasma assisted gas conversion or synthesis, not only in a GA but
28 also in other types of plasmas. For example, computer modeling is widely used to evaluate quantities
29 which are difficult to measure, and to identify the most important chemical reactions [26]-[32].
30 However, only a few papers in literature deal with modelling of a GA, typically applying a 1D
31 analytical model, such as the Elenbaas–Heller model [33] or the plasma string model [34] without a
32 detailed description of the chemical reactions occurring in the GA. Recently, a 2D non-quasi-neutral
33 model was presented to study the arc root movement in an argon GA [35]-[36]. Moreover, 3D quasi-
34 neutral models for a novel type of GA plasmatron [37] and a classical diverging electrode GA reactor
35 [38]-[39] were also reported. However, these models were all developed for argon. For a GA
36 operating in CO₂, the large number of species and related chemical reactions makes spatially
37 resolved models computationally expensive. That is why only a limited number of numerical studies
38 were reported so far on this subject, with only two papers for GA based CO₂ conversion published to

39 It is clear that more research is needed to fully exploit the capabilities of the GA for CO₂
40 conversion. In this paper, we therefore present a combined modeling and experimental study, based
41 – for the first time – on a 2D model. The aim of this study is not only to elucidate the underlying
42 mechanisms, but also – based on the obtained insights – to propose solutions on how to further
43 improve the performance of the GA for CO₂ conversion.

2. Experimental setup of the GA reactor

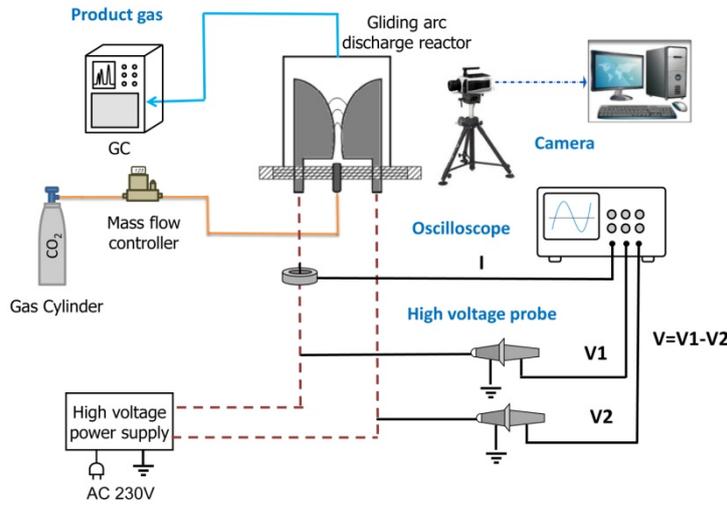


Figure 1 Schematic diagram of the GA experimental set-up

Figure 1 illustrates the experimental setup of the GA and surrounding measurement equipment. The GA reactor consists of two stainless steel semi-ellipsoidal electrodes with thickness of 2 mm (60 mm long and 18 mm wide) fixed in an insulating bracket and symmetrically placed on both sides of a gas nozzle with a diameter of 1.5 mm. The reactor is designed to facilitate easy electrode replacement, and the discharge gap between both electrodes, as well as the distance between the nozzle exit and electrode throat, is adjustable. Pure CO₂ gas was injected into the GA reactor and it pushes the arc plasma, which is initiated at the shortest gap between both electrodes, towards larger interelectrode distance until it extinguishes, and a new arc is created at the shortest gap. The plasma reactor was connected to a neon transformer (SIET, 230 V/10 kV, 50 Hz). The arc voltage was measured by a high voltage probe (Testec, TT-HVP 15 HF), while the arc current was recorded by a current monitor (Magnetlab, CT-E 0.5-BNC). All the electrical signals were sampled by a four-channel digital oscilloscope (Tektronix, MDO 3024). The arc dynamics are revealed by means of a digital high-speed camera (Phantom V.7.1) which can record up to 4,800 pictures per second using the full 800x600 pixel SR-CMOS imaging sensor array. The measurement technique was intensely optimized to fine-tune the best recording conditions. The frame rates to visualize the arc propagation and the exposure time of the detector to enhance the contrast between the arc and the reactor were investigated. The products of the CO₂ conversion after passing through the GA reactor were sampled when the plasma reaction has reached a stable condition, i.e., typically after 30 min. The gaseous products were analyzed by a gas chromatograph (Shimadzu, GC-2014) equipped with a thermal conductivity detector (TCD) and a flame ionization detector (FID). As we mention below, a standard case of 2.5 L/min and 40W is used to validate our model. Furthermore, the vertical distance between the nozzle exit and electrode throat was 2 mm and the shortest discharge gap between the two electrodes was also 2 mm.

The plasma power is calculated by integration of the arc voltage and current, as shown in Eq. (1).

$$P_{plasma} = 1/T \int_0^{t=T} V_{plasma} \times I_{plasma} dt \quad (1)$$

The conversion of CO₂, X_{CO_2} , is defined as:

$$X_{CO_2}(\%) = \frac{CO_{2(in)} - CO_{2(out)}}{CO_{2(in)}} \times 100\% \quad (2)$$

where $CO_{2(in)}$ and $CO_{2(out)}$ are the CO_2 signals without and with plasma, respectively. Since the method mentioned above does not account for the gas expansion due to CO_2 splitting, a correction factor is used, which is explained in the supplementary information of Ref [31].

In order to calculate the energy efficiency of CO_2 conversion, the specific energy input (SEI) in the plasma is defined as:

$$SEI \left(\frac{kJ}{L} \right) = \frac{\text{Plasma power (kW)}}{\text{Flow rate} \left(\frac{L_n}{min} \right)} \times 60 \left(\frac{s}{min} \right) \quad (3)$$

where the flow rate is expressed in L_n/min (liters normal per minute) with reference conditions at a temperature of $0^\circ C$ and a pressure of 1 atm.

The energy efficiency, η , is calculated as:

$$\eta(\%) = \frac{\Delta H_R \left(\frac{kJ}{mol} \right) \times X_{CO_2}(\%)}{SEI \left(\frac{kJ}{L} \right) \times 22.4 \left(\frac{L}{mol} \right)} \quad (4)$$

where ΔH_R is the reaction enthalpy of CO_2 splitting (i.e., 279.8 kJ/mol), X_{CO_2} is the amount of CO_2 converted, SEI is defined above and 22.4 L/mol is the molar volume at $0^\circ C$ and 1 atm.

The experiments were performed 4 times and they were reproducible within +/- 5% of the averaged values.

3. Description of the 2D plasma slab model

3.1 The GA reactor geometry

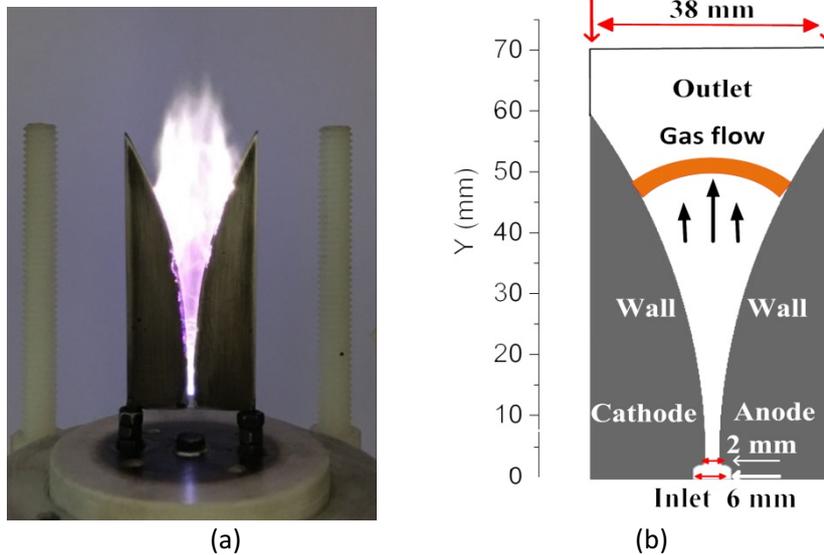


Figure 2 Photograph of the GA reactor (a) and schematic illustration of the geometry considered in the model (b).

The 2D fluid model that we developed applies to a Cartesian geometry, which allows to describe the gliding of a '2D arc', which is basically a finite plasma slab. The simulated geometry in the direction perpendicular to the simulation plane is assumed to be equal to the electrode thickness of 2 mm. Hence, the electrical current in the 2D model is obtained by integration of the current density

1 over the arc slab, which fits the experimental signal. Furthermore, the flow field is determined by
 2 taking into account a flow passing channel with a depth of 2 mm with the specified flow rate. In this
 3 way, the calculated gas velocity is similar to the experimental data when the vertical distance
 4 between the nozzle exit and electrode throat was 2 mm and the shortest discharge gap between the
 5 two electrodes was also 2 mm. Indeed, a rough estimation of the experimental gas velocity is
 6 obtained by examination of the arc displacement shown in successive high-speed photographs (see
 7 supporting information). In principle, a 3D model would be required to describe the GA behaviour in
 8 a realistic way, in view of the intrinsic 3D nature of the GA. However, a 3D model is very time
 9 consuming and it requires significant computer resources, especially when modelling a CO₂ plasma
 10 with complicated plasma chemistry. Furthermore, previous work for an argon GA [38] has shown
 11 that the results of a 2D model compare well with those of a 3D model, and can thus be used for a
 12 better understanding of the GA basic characteristics. The total width and height of the model
 13 geometry, including the region outside the electrodes where the gas can flow without passing
 14 through the arc, is 38 mm and 70 mm, respectively.

15 3.2 CO₂ plasma chemistry and treatment of the vibrational levels

16 The chemistry set is based on the full chemistry set developed by Kozák and Bogaerts [26]-[27]
 17 with a 0D model, but reduced to include only the most important species and processes. In this way,
 18 we can avoid excessive calculation times in this 2D model, but we still account for the vibrational
 19 kinetics, which is crucial for describing CO₂ conversion in a GA plasma reactor [41]. The list of species
 20 considered in the model is shown in table 1. These species include various neutral molecules in the
 21 ground state, as well as in various electronically and vibrationally excited levels, a number of radicals,
 22 positive and negative ions, and the electrons. In the full model of Kozák and Bogaerts [26]-[27], 25
 23 CO₂ vibrational levels (i.e., 4 effective levels of the symmetric modes and 21 levels of the asymmetric
 24 stretch mode, up to the dissociation limit) were taken into account. However, to further reduce the
 25 calculation time, which is needed to implement this chemistry in a 2D model, Berthelot and Bogaerts
 26 [42] developed a level lumping method, which groups the 21 asymmetric stretch mode vibrational
 27 levels into a number of lumped levels, without loss of essential information. We applied this level
 28 lumping method in [41] for a 1D gliding arc model, and we illustrated that lumping the 21 levels into
 29 3 groups can reproduce the plasma properties, the vibrational distribution function (VDF) and the
 30 CO₂ conversion very well. Therefore, we adopt here the same level lumping method with 3 groups
 31 for the asymmetric stretch mode, with each group including 7 vibrational levels (group 1: CO₂[v₁-v₇],
 32 group 2: CO₂[v₈-v₁₄], group 3: CO₂[v₁₅-v₂₁]). The species number density of each level within one
 33 group can be determined following the method described in [41-42]. Besides, we also take into
 34 account the 4 effective levels of the symmetric modes (CO₂[v_a] – CO₂[v_d]), 1 electronically excited
 35 level of CO₂ (CO₂[e]), and 3 vibrational levels of O₂ (O₂[v1] - O₂[v3]), as indicated in table 1.

36 Table 1 Overview of the plasma species included in the model.

Neutral ground state species	CO ₂ , CO, C, O ₂ , O
Neutral excited states	CO ₂ [v _a], CO ₂ [v _b], CO ₂ [v _c], CO ₂ [v _d], CO ₂ (v ₁ -v ₇), CO ₂ [v ₈ -v ₁₄], CO ₂ [v ₁₅ -v ₂₁], CO ₂ [e], O ₂ [v1], O ₂ [v2], O ₂ [v3]
Charged species	CO ₂ ⁺ , O ₂ ⁺ , CO ₃ ⁻ , O ⁻ , O ₂ ⁻ , e ⁻

37 All these species undergo a large number of chemical reactions, i.e., electron impact collisions
 38 with neutral species, leading to excitation, ionization, dissociation and electron attachment,

1 electron-ion recombination reactions, as well as many heavy-particle chemical reactions (i.e., ion-ion,
2 ion-neutral and neutral-neutral reactions). We pay special attention to the reactions of the
3 vibrational levels, i.e., electron impact vibrational excitation, and vibrational energy exchange upon
4 collision with ground state species or other vibrationally excited levels (i.e., so-called vibrational-
5 translational (VT) and vibrational-vibrational (VV) relaxation, respectively). Moreover, the same
6 chemical reactions as for the ground state species are carefully included for the vibrational levels as
7 well, because the vibrational energy can help overcome the activation energy barrier of the reactions
8 and thus increase the reaction rate of CO₂ splitting. The chemical reactions, the corresponding rate
9 coefficients and the references where these data were adopted from, are listed in our previous work
10 [41].

11 **3.3 System of governing equations and boundary conditions**

12 The model calculates the densities of all the plasma species, the electron temperature and gas
13 temperature and the electric field in the GA, as well as the gas flow profile. We assume electrical
14 neutrality in the plasma, because the sheath is not considered in our model. This assumption has no
15 significant influence on the arc column [39]. The species densities and the electron mean energy are
16 calculated with continuity equations based on transport and on production and loss terms defined by
17 the chemical reactions (and by Joule heating for the electron energy). The species transport is based
18 on drift in the electric field and diffusion due to concentration gradients. As we assume electrical
19 neutrality in the arc plasma, the ambipolar electric field is calculated from the charged species
20 densities. The gas heat transfer equation is solved for the gas translational temperature, and finally,
21 the neutral gas flow, which is responsible for the arc displacement, is described by the Navier-Stokes
22 equations, providing a solution for the mass density and the mass-averaged velocity. The Navier-
23 Stokes equations are first solved separately, and subsequently, the obtained velocity distribution is
24 used as input data in the other equations, describing the plasma behavior and the gas heating. The
25 equations solved, as well as the corresponding boundary conditions, are explained in detail in the
26 supporting information. Finally, the external circuit and the power supply need to be specified in the
27 simulation. The source voltage has a sinus shape, $V_{\text{source}} = 7200\sin(2\pi 50t + 0.50)$ V, and a resistance
28 of 60 k Ω is used to limit the discharge current; it provides a total arc discharge power of 40 W, which
29 is similar to the typical experimental value at a gas flow rate of 2.5 L/min.

30 The equations are solved by means of the COMSOL Multiphysics software [43], a commercial
31 finite element software designed for solving problems of multi-physics. As initial values we assume
32 that the concentrations of CO₂ in the ground state and in the various excited levels follow a
33 Maxwellian distribution at room temperature.

34 **4. Results and discussion**

35 In section 4.1 we will first validate our model by comparing our calculated values with
36 experimental data for the electron number density (which is one of the most important plasma
37 properties), as well as for the CO₂ conversion and corresponding energy efficiency. Subsequently, in
38 section 4.2 we will benchmark our results for the CO₂ conversion and energy efficiency to the
39 classical thermal conversion process and to other plasma-based technologies for CO₂ conversion
40 reported in literature. This allows us to provide a clear overview of the capabilities of the GA for CO₂
41 conversion, as well as its limitations, for which we should propose some process modifications, to
42 further improve the results. In order to achieve this, we need a better insight in the typical discharge
43 characteristics, as calculated by the model, which will be presented in section 4.3. Furthermore, we

1 will also perform a chemical kinetics analysis in section 4.4, to elucidate the role of various plasma
 2 species and their reactions in the GA based CO₂ conversion. Finally, based on the revealed discharge
 3 properties and the obtained plasma chemistry, we will propose in section 4.5 some solutions on how
 4 to further improve the CO₂ conversion and the energy efficiency by the GA.

5 4.1 Experimental validation of the model

6 Table 2 Comparison of our calculated values for electron number density, CO₂ conversion and
 7 energy efficiency, with the experimental data, at a gas flow rate of 2.5 L/min and a discharge power
 8 of 40 W.

Results	Electron number density	Conversion	Energy efficiency
Calculation	10^{18} - 10^{19} m ⁻³	2.78 %	32.8 %
Experiment	2.6×10^{18} m ⁻³	2.90 %	34.3 %
Experimental error	4.9%	4.3%	4.6%

9 In table 2 we compare our calculated results for the electron number density, CO₂ conversion
 10 and corresponding energy efficiency with the corresponding measured values, at a typical
 11 experimental gas flow rate of 2.5 L/min and a discharge power of 40 W.

12 The experimental electron number density is obtained from the electrical characteristics and the
 13 high speed camera images, as follows. During the propagating phase of the GA, the average
 14 experimental voltage drop across the arcs is $V \approx 1.20$ kV with an average current of $I \approx 0.06$ A (see
 15 figure S1 of the supporting information), leading to an average arc impedance $\langle R \rangle = V/I \approx 20$ kΩ. The
 16 radius of the arc ($\lambda \approx 1$ mm) and the average length ($\langle w \rangle \approx 15$ mm) are obtained by the high speed
 17 camera recordings (see figure S2 of the supporting information). With this information, we can
 18 calculate the average arc electrical conductivity, σ , as

$$19 \quad \sigma = \frac{\langle w \rangle}{\langle R \rangle \pi \lambda^2} \quad (5)$$

20 yielding $\sigma \approx 0.24$ S/m. The conductivity can be related to the electron density through the electron
 21 mobility, μ_e , using:

$$22 \quad \langle n_e \rangle = \frac{\sigma}{e \mu_e} \quad (6)$$

23 With e the electron charge. Using a time averaged gas temperature of 2400 K and an electron
 24 temperature of 1.7 eV, as obtained from our model (see section 4.3), we calculated $\mu_e =$
 25 0.56 m²/V/S by means of a Boltzmann equation solver BOLSIG+ [44]. Hence, formula (6) gives an
 26 estimate of the time and spatially averaged electron number density, $\langle n_e \rangle \approx 2.6 \times 10^{18}$ m⁻³. Our
 27 calculations predict the maximum electron number density in the discharge channel to be around
 28 10^{19} m⁻³ (see section 4.3). Considering the non-uniform distribution within the discharge channel, we
 29 can obtain a spatially averaged value of the electron number density within the range 10^{18} m⁻³- 10^{19}
 30 m⁻³, indicating a reasonable agreement between the calculated and measured values.

31 The calculated conversion of CO₂, $X_{CO_2}^C$, is determined as:

$$32 \quad X_{CO_2}^C = \frac{\int r_{CO_2} dv \int dt}{\int Q_{CO_2(in)} dt} \times 100\% = \frac{l_0 \int r_{CO_2} ds \int dt}{\int Q_{CO_2(in)} dt} \times 100\% \quad (7)$$

1 where $Q_{CO_2(in)}$ is the particle flow rate of CO₂ entering the reactor per second (in s⁻¹), r_{CO_2} is the net
 2 splitting rate of CO₂ inside the arc (in m⁻³s⁻¹), and $l_0 = 2$ mm, is the thickness of the GA reactor (see
 3 below).

4 The particle flow rate of CO₂, $Q_{CO_2(in)}$, represents the total number of CO₂ molecules flowing
 5 into the reactor per second, and is obtained as follows:

$$6 \quad Q_{CO_2(in)} \left(\frac{1}{s} \right) = \frac{Q_n \left(\frac{L_n}{min} \right) \times 0.001 \left(\frac{m^3}{L_n} \right) \times \frac{1}{60} \left(\frac{min}{s} \right) \times P_0 (Pa)}{k \left(\frac{J}{K} \right) \times T_0 (K)} \times 100\% \quad (8)$$

7 where k is the Boltzmann constant, Q_n is the gas flow rate at the standard temperature $T_0 = 273$ K
 8 and pressure $P_0 = 101325$ Pa.

9 The net splitting rate of CO₂, r_{CO_2} in m⁻³s⁻¹, represents the net number of dissociated CO₂
 10 molecules per volume and per second, and is obtained by taking into account all the chemical
 11 reactions, leading to destruction (when a positive value) or formation (when negative) of CO₂
 12 molecules. In order to determine the total conversion of CO₂, as shown in equation (6), the net
 13 splitting rate of CO₂, r_{CO_2} , is integrated spatially over the whole reactor and temporally over the
 14 whole gliding cycle. Because of the prohibitively long computation time in a 3D model, a 2D plasma
 15 slab model is used, assuming that the distribution of plasma parameters in the direction
 16 perpendicular to the simulation plane (see figure 2b) is uniform. As a result, the arc is not a “wire”
 17 but a “slab” with a length l_0 in the direction perpendicular to the simulation plane. We assume l_0 is
 18 equal to the thickness of the GA reactor, i.e., 2 mm. Thus the total conversion of CO₂ in the 2D model
 19 is obtained by the integration of the net splitting rate of CO₂ r_{CO_2} over the arc slab with $l_0 = 2$ mm.

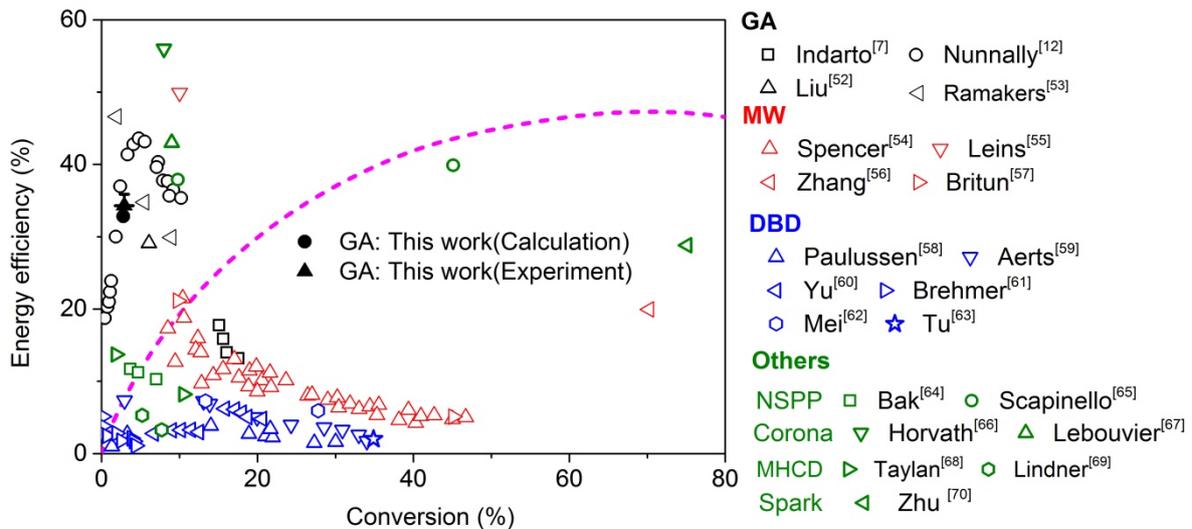
20 Our calculated conversion and energy efficiency of CO₂, at a gas flow rate of 2.5 L/min and a
 21 discharge power of 40 W, are 2.78 % and 32.8 %, respectively, which is also in satisfactory agreement
 22 with the experimental values of 2.90 % and 34.3 %. The comparison of these three key parameters
 23 indicates that our model most probably can provide a realistic picture of the plasma chemistry.

24 Comparison of other plasma characteristics, such as the electron temperature or gas
 25 temperature, was not possible, as the latter properties could not be determined in our experimental
 26 setup, and are also not available in literature for a pure CO₂ GA. This is probably because optical
 27 emission spectrometry is not suitable here, as there are no proper spectral lines that can be used.
 28 However, our calculated values for electron temperature (up to 1.7 eV) and gas temperature (up to
 29 around 2700 K) are comparable with experimental data from literature, for GA reactors using other
 30 molecular gases (nitrogen and air) [45]-[47], as well as for gaseous mixtures containing CO₂ [48]- [49].
 31 For example, Wu et al. [48] measured values for the electron excitation temperature of
 32 approximately 1.1-1.7 eV, using a rotating GA reactor for a mixture of CH₄/CO₂. Moreover, in a non-
 33 equilibrium GA "tornado" discharge using CO₂ doped with 1% N₂, the rotational gas temperature was
 34 determined to be 2700K ± 50 K [49].

35 We can only compare here the calculated and experimental data at a gas flow rate of 2.5 L/min
 36 and discharge power of 40 W, because at these conditions the arc was observed to glide smoothly
 37 along the electrodes. Indeed, at higher gas flow rates, a phenomenon of back-breakdown occurs,
 38 affecting the arc gliding process (see further). These back-breakdown events cannot self-consistently
 39 be captured by the model, because this behaviour is mostly stochastic by nature and the arc
 40 instabilities are not well defined. Therefore, we would need to make some assumptions in the model

1 on the number of back-breakdown events, and depending on the values assumed for the back-
 2 breakdown frequency, we would always be able to obtain good agreement with the experiments.
 3 Hence, we lose the real validation possibility at higher gas flow rates. Therefore, we could only
 4 validate the model at a gas flow rate of 2.5 L/min and a discharge power of 40 W, where our high
 5 speed camera did not record any back-breakdown events. However, in section 4.5, we will assess the
 6 effect of a different number of back-breakdown events on the calculated conversion and energy
 7 efficiency, which can in principle be correlated with different values of gas flow rate and discharge
 8 power.

9 4.2 Comparison of our results with other plasma systems from literature



10 **Figure 3** Energy efficiency vs CO₂ conversion, obtained in our experiments and calculations (black
 11 solid symbols), and comparison with other GA results from literature (black open symbols), as well as
 12 with results from other types of plasma reactors used for CO₂ conversion, operating at atmospheric
 13 pressure. The thermal conversion limit is also indicated.

14
 15 In figure 3, we compare our results for the energy efficiency vs CO₂ conversion with data
 16 obtained from literature for CO₂ splitting, in other GA discharges[5],[6],[50],**Error! Reference source**
 17 **not found.**, as well as in other types of plasma reactors, such as microwave (MW) plasma [51]-[55],
 18 dielectric barrier discharge (DBD) [56]-[61], nano-second pulsed plasma (NSPP) [62]-[63], corona
 19 discharge [64]-[65], micro hollow cathode discharge (MHCD) [66]-[67] and spark discharge [68]. We
 20 can conclude that in terms of energy efficiency, the GA plasma is very promising, similar to the
 21 corona discharge [64]-[65]. It should be mentioned that for MW plasmas some higher energy
 22 efficiencies (i.e., up to 80 and 90%) were obtained in literature by Rusanov et al.[69] and Asisov et
 23 al.[70], respectively. However, their MW plasma reactors were operating at a reduced pressure of
 24 0.06 - 0.26 atm and 0.05 – 0.2 atm, respectively, and thus they need vacuum equipment. This makes
 25 it more difficult to be applied on industrial scale, and the energy cost of the pumping system should
 26 also be included when calculating the energy consumption. Bongers et al. recently obtained values
 27 up to 50%, when applying a reverse vortex gas flow [73], but again these experiments were
 28 conducted at reduced pressures of 150 – 600 mbar (0.15 – 0.60 atm). In order to allow a fair
 29 comparison, we therefore only present results in figure 3, obtained at atmospheric pressure. When
 30 the MW discharge is operating at atmospheric pressure, the reported energy efficiency dramatically
 31 drops to values of about 5 - 20 % [51]-[55].

1 If we compare our results with those obtained in other GA reactors from literature, it is
2 important to explain that there exist roughly two different reactor designs. The classical GA reactor,
3 which is used in this study, typically consists of two plane diverging electrodes between which the gas
4 flows. In contrast, recently a three-dimensional GA reactor, consisting of cylindrical electrodes with
5 tangential gas inlet, leading to a vortex gas flow configuration, has been developed, also called GA
6 plasmatron (GAP) [19]. Indarto et al. [5] applied a classical GA configuration, like in our case, and they
7 obtained a highest energy efficiency of around 17%, which is much lower than our current work. On
8 the other hand, Nunnally et al. [6], Liu et al. [50] and Ramakers et al. **Error! Reference source not**
9 **found.** used a vortex flow GAP, which can reach a somewhat higher conversion and energy efficiency.
10 This reactor design is indeed very promising, because it can be more easily implemented in industry
11 and the specific gas flow configuration ensures the gas treatment to be more uniform. This indicates
12 that a better design of the classical GA reactor, to enhance the treated gas volume, would improve
13 the conversion performance, as will be discussed in detail in section 4.5 below. However, in general
14 we can deduce from figure 3 that the GA plasma shows a very good performance with a relatively
15 high energy efficiency. This is because the energy efficient vibrational excitation processes are
16 favoured, as will be revealed in section 4.4 below.

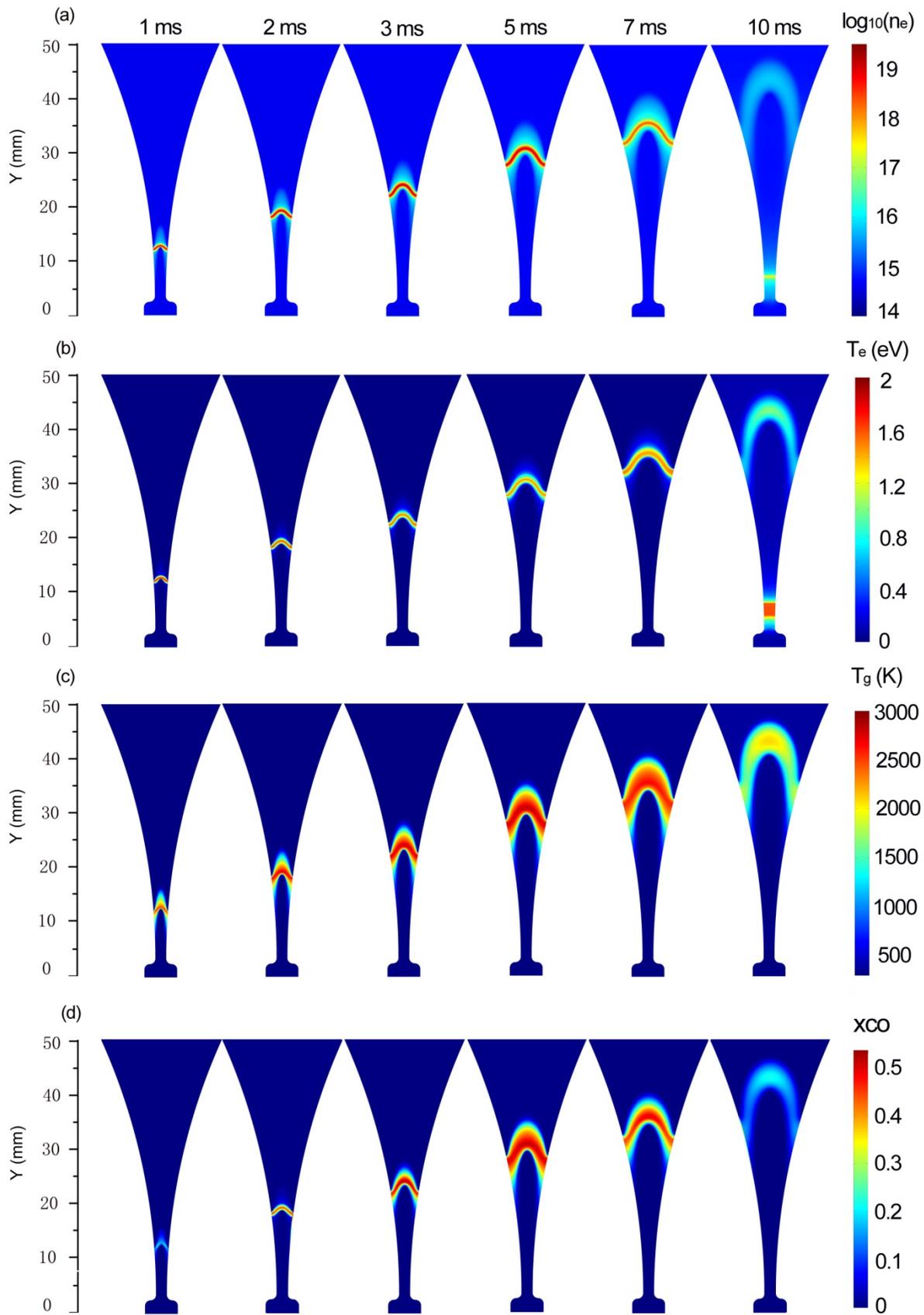
17 It is obvious from figure 3 that a DBD plasma [56]-[61] has a reasonable conversion but a quite
18 low energy efficiency. This is due to the non-ideal operating conditions, as the electron temperature
19 is typically higher than in a GA (or MW) plasma [19],[72], and the mechanism of CO₂ conversion
20 involves charged and electronically excited species, and thus it is limited by the high energy cost for
21 the formation of these species. The same applies for the nano-second pulsed plasma (NSPP) [62]-[63]
22 which also has a rather low energy efficiency. The process capability of the micro hollow cathode
23 discharge (MHCD) [66]-[67] is very limited due to its very small volume. Therefore, it generally also
24 exhibits a relatively low energy efficiency. The spark discharge [68] has a very high conversion,
25 because of the very high energy consumption. The energy efficiency is also quite high, but it is lower
26 than the thermal conversion process. This may be attributed to the fact that most of the energy is
27 spent on the gas heating and the energy exchange with the surroundings. In general, we can
28 conclude that the energy efficiency in our GA reactor at atmospheric pressure is better than the DBD
29 plasma, microwave plasma, nano-second pulsed plasma and micro hollow cathode discharge plasma,
30 and comparable to the corona discharge [64]-[65].

31 Finally, we also benchmark our results for the GA based CO₂ conversion to the pure thermal
32 conversion process (see the calculation method for the latter in the supporting information). It is
33 clear that the CO₂ conversion in our GA proceeds more energy efficient than pure thermal
34 conversion. This is because the energy in the thermal conversion is distributed over all degrees of
35 freedom based on the equipartition principle of energy, and thus it is especially spent on gas heating
36 rather than on CO₂ dissociation reactions. In contrast, our GA clearly operates in non-equilibrium
37 conditions, as the electrons have a much higher temperature than the gas itself (see our calculation
38 results in section 4.3 below). These highly energetic electrons induce different chemical reactions,
39 which normally do not occur at the considered gas temperate in case of equilibrium conditions.

40 In spite of the reasonable results obtained already by the gliding arc, the conversion should still
41 be further improved, while maintaining the high energy efficiency. More specifically, if this low
42 conversion could not be further improved, it would imply the need for operating in a recycle mode,
43 which would make the system highly non-effective.

1 **4.3 Typical GA discharge characteristics**

2 In order to understand the time behavior of the plasma characteristics in the CO₂ GA, we plot in
3 figure 4 the electron number density, electron temperature, gas temperature, as well as of the CO
4 molar fraction distribution, at different moments in time, for a gas flow rate of 2.5 L/min and a
5 discharge power of 40 W.



1
 2 **Figure 4** Time evolution of the electron number density (in m^{-3}), electron temperature, gas
 3 temperature and CO molar fraction distribution, at a gas flow rate of 2.5 L/min and a discharge
 4 power of 40 W.

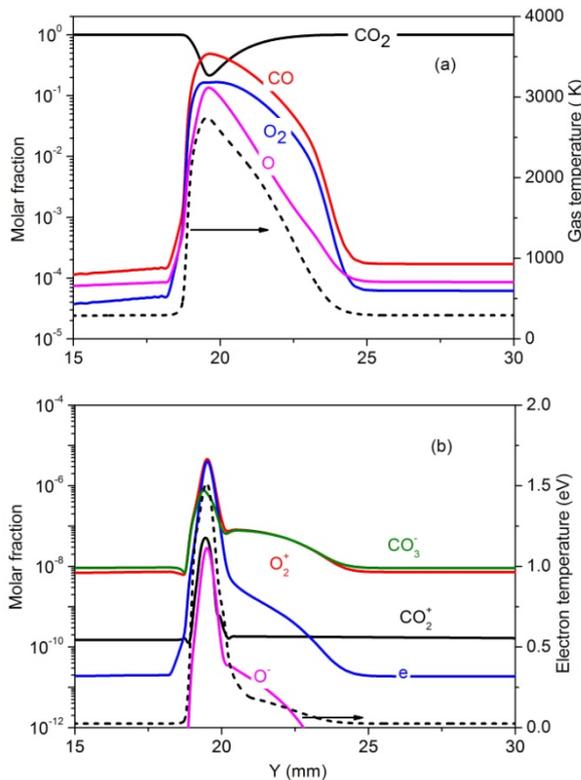
1 The results are plotted starting from $t = 1$ ms. At $t = 0$ ms, the source voltage is larger than the
2 critical breakdown voltage with a shortest gap separation of 2 mm. The discharge ignition takes place,
3 because of a positive value of the net electron generation, yielding an abrupt increase of the electron
4 number density during the electrical breakdown. Once the conducting channel is established, the arc
5 travels along the electrodes as a result of the gas flow drag. Since the gas velocity has a maximum
6 value at the discharge axis and gradually decreases to zero at the electrode surface, the arc root
7 moves at a much slower velocity compared to the arc body. Thus, the arc gradually begins to bend
8 due to the gas blast. The maximum electron number density also increases due to the rising voltage
9 and hence discharge current (see figure S1 in the supporting information), till a peak value is reached
10 at 3.5 ms (see figure 4(a)). At later times, the discharge current drop, and consequently, the electron
11 density follows the same trend till zero at $t = 8.5$ ms, when the applied voltage reaches zero (see
12 figure S1). The GA gradually extinguishes and enters a relaxation stage, where the voltage is small
13 and not enough to sustain the GA discharge. Thus, there is a decaying residual low density plasma
14 moving downstream with the gas flow (see figure 4(a)). Shortly after $t = 8.5$ ms, the applied voltage
15 of the alternating current (AC) power source changes its polarity (see figure S1 of the supporting
16 information) and reaches again the critical breakdown voltage at the narrowest electrode gap
17 separation of 2 mm, where a restrike occurs by establishing a new conducting channel. It should be
18 noted that the re-ignition of the GA does not exactly take place at the shortest gap separation ($Y =$
19 2.5 mm), but at $Y = 7.5$ mm. This is because the local electric field at $Y = 7.5$ mm first reaches the
20 critical breakdown field. This is in good agreement with our experiments, recorded by the digital
21 camera (see figure S2 of the supporting information).

22 The rise and drop in electron number density during one GA discharge cycle results in an
23 enhanced and reduced Joule heating effect before and after $t = 3.5$ ms, respectively. The Joule
24 heating refers to the process by which the passage of an electric current through a conductive
25 medium produces heat and causes heating of the electrons. Correspondingly, the electron
26 temperature first increases and then decreases (see figure 4(b)). After $t = 8.5$ ms, the electron
27 temperature of the residual GA channel continuously decreases, because the electron number
28 density and the electric energy stored in the channel decay very rapidly. Subsequently, the extremely
29 large reverse polarity voltage imposed across the electrodes at the shortest electrode gap leads again
30 to an increase of the electron temperature and hence a subsequent breakdown at the new position
31 of $Y = 7.5$ mm (see figure 4(b)).

32 Once the discharge is ignited, the electrons cause vibrational excitation of CO_2 , and the energy
33 stored in the vibrationally excited states will partially be transferred to the gas by vibrational-
34 translational (V-T) relaxation. Indeed, at atmospheric pressure, the typical characteristic time for V-T
35 relaxation in CO_2 is very short (around 10^{-5} s). As a result, the gas temperature also rises as a function
36 of time, reaching a maximum value of about 2700 K at around $t = 3.5$ ms, when the applied source
37 voltage ($V_{\text{source}} = 7200\sin(2\pi 50t + 0.50)$) and the discharge current reach their maximum (see figure
38 4(c)). Subsequently, the gas temperature in the arc channel gradually decreases to around 2000 K
39 when a new cycle starts at $t = 10$ ms, because the discharge power decays rapidly in the relaxation
40 stage from 8.5 ms to 10 ms.

41 The CO molar fraction is obviously equal to zero before the arc is formed, but it starts increasing
42 gradually as a function of time, when the voltage and hence the discharge current in the arc rise, up
43 to a value of 0.55 at $t = 3.5$ ms, indicating that CO_2 is gradually converted into CO. At later times, the

1 discharge current and hence the discharge power start to drop, so the CO molar fraction within the
 2 arc channel gradually decreases until the arc is extinguished. This is caused by recombination of CO
 3 and O into CO₂. Furthermore, new CO₂ gas will continuously be transported into the arc channel by
 4 both diffusion and convection, while the dissociation products will leave the discharge channel by the
 5 same transport mechanisms. This leads to a reduction of the maximum local CO molar fraction, as is
 6 clearly indicated in figure 4 (d). Note that the overall CO₂ conversion is much lower than the local
 7 conversion of 80%, which corresponds to the maximum CO molar fraction of 0.55 (and CO₂ molar
 8 fraction of 0.2; see below). This is because the overall CO₂ conversion is calculated for the entire gas
 9 passing through the reactor, integrated over the time of one GA cycle (i.e., 10 ms), and thus not only
 10 for the fraction of gas passing through the active arc channel at a certain moment in time.



11
 12 **Figure 5** 1D distribution of the molar fractions of the neutral species (a) and the charged species (b)
 13 as a function of axial distance on the symmetry plane, at a time instant of 2.5 ms. The gas
 14 temperature and electron temperature are also plotted in dashed lines in (a) and (b), respectively.

15 The molar fractions of the major neutral and charged species occurring in the CO₂ GA are plotted
 16 as a function of Y position in figure 5, at a time instant of 2.5 ms, and at the same conditions as in
 17 figure 4. It is clear that CO₂ is the major component in the plasma, except at the centre of the arc,
 18 where the molar fraction of CO₂ (around 0.2) is lower than the fraction of CO (around 0.5), and
 19 comparable to the molar fractions of O₂ (0.16) and O (0.14). This indicates that the majority of CO₂ is
 20 split here into CO and O₂ as well as O atoms. Moreover, part of the O atoms have recombined into O₂
 21 molecules, indicating a higher decay rate of the O molar fraction than that of O₂. The molar fractions
 22 of CO, O and O₂ drop quickly when moving towards the outer part of the arc, indicating that most of
 23 the CO₂ splitting takes place in the centre of the arc.

24 The CO₂ conversion can be further enhanced when applying a higher power, however, even at
 25 100 W, the local molar fraction of CO₂ drops to extremely low values and the local conversion in the

1 GA reaches almost 100 %. This limits the further improvement of GA based CO₂ conversion.
 2 Therefore, the conversion can only be further enhanced if we can provide more CO₂ into the arc
 3 centre, while at the same time remove the dissociation products (CO and O₂) out of the arc centre.
 4 This will be further discussed in detail in section 4.5.

5 The molar fractions of the various charged species are at maximum 10⁻⁵, even in the arc centre,
 6 and they clearly drop upon larger distance from the centre of the arc. Also the electron molar fraction
 7 is at maximum 10⁻⁵, indicating that the CO₂ plasma is only weakly ionized, even in the centre of the
 8 arc. The major positive ions are the O₂⁺ ions, while the CO₃⁻ ions are the major negative ions, and
 9 they are even more important (although still with very low molar fractions) than the electrons, except
 10 in the centre of the arc. These trends are in agreement with our previous findings obtained by a 1D
 11 cylindrical discharge model, despite the considerable number of approximations adopted there [41].

12 The gas temperature and electron temperature are also plotted in figure 5. They both reach their
 13 maximum in the centre of the arc, as is logical, and they drop significantly as a function of position
 14 from the arc centre. The electron temperature reaches a maximum of 1.5 eV (or 17,400 K) in the
 15 centre of the arc at the time instant of 2.5 ms, but it drops significantly as a function of rising distance
 16 from the arc centre in the first 0.5 mm, followed by a slower decay to thermal values at a distance of
 17 about 1.0 mm from the centre. The gas temperature is at maximum about 2700 K in the centre of the
 18 arc. From the comparison between these temperatures, it is clear that the gliding arc is far from
 19 thermal equilibrium, as the electron temperature is about 6 times higher than the gas temperature.
 20 As mentioned in section 4.1 above, a gas temperature up to around 2700 K and an electron
 21 temperature up to 1.5 eV correspond well to experimental data found in literature for low current
 22 atmospheric pressure GA discharges, although it should be mentioned that it is not easy to compare
 23 different GA setups with different reactor geometries and discharge conditions.

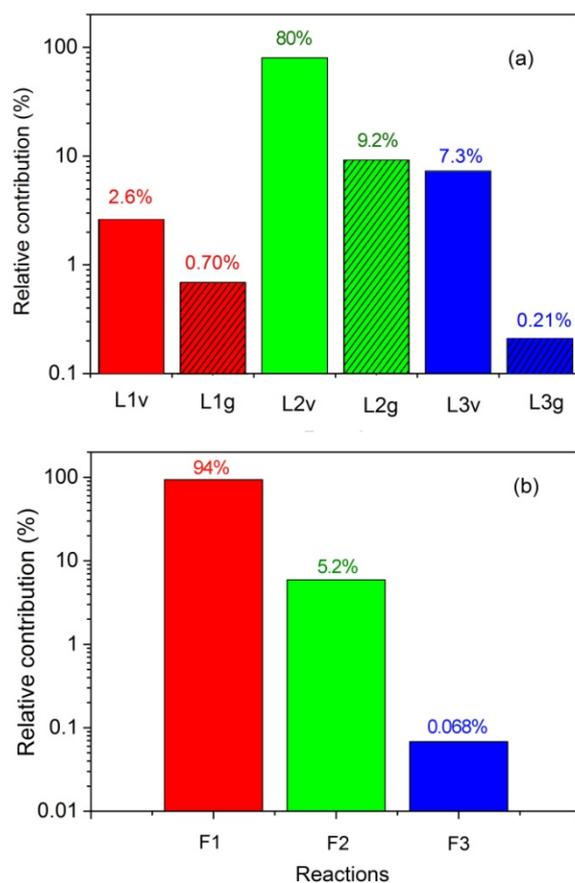
24 4.4 CO₂ conversion mechanisms in the GA

25 In order to evaluate which mechanisms are the most important for the CO₂ splitting in the GA
 26 plasma, and how they can eventually be further improved, we investigated the dominant reaction
 27 pathways for the formation and loss of CO₂ for the same conditions as in figure 4. The reactions are
 28 listed in table 3 and their relative contributions to the overall CO₂ loss and formation are presented in
 29 figure 6. This kinetic analysis was performed by looking at the time and volume integrated rates of
 30 the various processes for a complete gliding cycle of 10 ms. In the supporting information, we also
 31 plot the temporal evolution of the most important loss and formation rates of CO₂, obtained by
 32 integrating the reaction rates over the entire reactor (see figure S4).

33 Table 3 Dominant CO₂ loss and formation reactions.

Process	Loss reaction	Process	Formation reaction
L1v	$e + CO_2(v) \rightarrow e + CO + O$	F1	$CO + O_2 \rightarrow CO_2 + O^{(a)}$
L1g	$e + CO_2(g) \rightarrow e + CO + O$		
L2v	$CO_2(v) + O \rightarrow CO + O_2$	F2	$CO + O + M \rightarrow CO_2 + M$
L2g	$CO_2(g) + O \rightarrow CO + O_2$		
L3v	$CO_2(v) + M \rightarrow CO + O + M$	F3	$CO + O^- \rightarrow e + CO_2$
L3g	$CO_2(g) + M \rightarrow CO + O + M$		
L4v	$CO_2(v) + O^- + M \rightarrow CO_3^- + M$		
L4g	$CO_2(g) + O^- + M \rightarrow CO_3^- + M$		

34 (a) O₂ represents the sum of the ground state and the vibrational states of molecular oxygen.



1
2 **Figure 6** Relative contributions of the most important processes for CO₂ loss (a) and formation (b).
3 The reaction numbers in the x-axis correspond to the numbers in table 3. Note that only the three
4 main loss processes are illustrated, as the fourth process (L4v, L4g) contributes for less than 0.1 %.

5 The most important process for CO₂ loss is the dissociation of vibrationally excited states of CO₂
6 upon collision with O atoms (L2v) with a relative contribution of about 80 %. The same process, but
7 upon collision of ground state CO₂ with O atoms (L2g) has a relative contribution of 9.2 %.
8 Furthermore, the dissociation of vibrationally excited states of CO₂ upon collision with any neutral
9 species (M) also contributes for 7.3 % (L3v). The relative contribution of the same process, but
10 starting from ground state CO₂, is only 0.21 % (L3g). Besides, electron impact dissociation from the
11 CO₂ vibrational levels (L1v) and from the CO₂ ground state (L1g) contribute for 2.6 % and 0.70 %,
12 respectively. Compared with the electron impact dissociation reactions, the neutral reactions upon
13 collision with O atoms have a lower energy requirement [19] and hence are more energy efficient.
14 Note that reactions L2v and L2g are actually follow-up reactions of reactions L1v and L1g, as the O
15 atom that reacts in reactions L2v and L2g is the result of CO₂ splitting, either by reactions L1v and L1g,
16 or reactions L3v and L3g. Nevertheless, once the first O atoms are formed upon CO₂ splitting, the
17 reactions L2v and L2g can occur in parallel to these other reactions, and thus we can consider them
18 separately in this analysis.

19 Our calculation results reveal that the CO₂ dissociation mainly proceeds from the vibrationally
20 excited levels of CO₂. The latter provide more energy efficient dissociation, because the vibrational
21 energy can help overcome the activation energy barrier of the reaction and thus increase the
22 reaction rate constant [26]-[27]. This is consistent with experimental investigations in literature.
23 Indeed, experimental work for both a diverging electrodes gliding arc reactor [5] and a gliding arc

1 plasmatron [6] shows that the presence of a very small quantity of water added into CO₂ greatly
2 reduces the power efficiency compared with pure CO₂ at atmospheric pressure. This is explained by
3 the fact that water can significantly reduce the vibrational excitation of CO₂ molecules, because the
4 energy is absorbed and quickly lost by water. Based on this, Nunnally et al. [6] concluded that non-
5 equilibrium vibrational excitation plays the major role during CO₂ dissociation in a gliding arc.

6 Additionally, there exist measurements in the literature, demonstrating that the vibrational
7 temperature in the gliding arc is higher than the gas temperature, even at atmospheric pressure,
8 although we cannot validate our model by direct comparison, as experimental data for the
9 vibrational temperature in pure CO₂ in classical gliding arc reactors do not yet exist. However, in a
10 non-equilibrium gliding arc "tornado" discharge using CO₂ doped with 1% N₂ at a flow rate of 10 lpm
11 and a power of 200 W, Nunnally et al. [49] estimated the vibrational temperature to be
12 approximately 6000 K at atmospheric pressure, by comparing the theoretical and experimentally
13 measured spectra for the N₂ system, and this value is much higher than the reported rotational gas
14 temperature of 2700K ± 50K. Therefore, these experimental results support our modelling results.

15 Some of the reactions plotted in figure 6(a) also occur in the opposite direction, hence, besides
16 dissociation of CO₂, the recombination of CO with O₂, O and O⁻ ions also takes place in the GA, giving
17 rise to the formation of CO₂ again and yielding a lower net conversion of CO₂. The recombination
18 reaction of CO with O₂ molecules (F1, i.e., the opposite of L2) is the predominant production process
19 of CO₂, with a relative contribution to the overall CO₂ formation amounting to 94 %. The
20 recombination reaction of CO with O atoms (F2, i.e., the opposite of L3) has a relative contribution of
21 5.2 %, while the recombination of CO with O⁻ ions (F3) only contributes for 0.068 %. Other reactions
22 play a negligible role towards CO₂ formation (< 0.05 %).

23 Note that the reverse reactions, especially the recombination of CO with O₂ molecules, have
24 only slightly lower rates than the rates of the most important loss processes, as depicted in figure S4
25 in the supporting information. Therefore, these reactions have a detrimental effect on the overall
26 CO₂ conversion. Indeed, when the rates of these reactions would become even larger, they would
27 inhibit further CO₂ dissociation. This happens when a considerable fraction of the CO₂ molecules is
28 already converted into CO and O/O₂, and especially at high gas temperature in the arc. When
29 comparing the total loss of CO₂, integrated over the entire arc and the whole gliding cycle, with the
30 total formation of CO₂, we obtain values of 3.8 x 10⁻¹⁸ vs 3.5 x 10⁻¹⁸ at the conditions under study. Thus,
31 it is clear that about 92 % of the CO₂ converted in the GA, will be formed again, so the net conversion
32 of CO₂ into CO is much smaller than the initial loss of CO₂. Therefore, the recombination of CO with
33 O₂ back into CO₂ is clearly a limiting factor, which affects the further improvement of GA based CO₂
34 conversion and its energy efficiency. This will be discussed in the next section.

35 **4.5 How to improve the CO₂ conversion and energy efficiency in the GA ?**

36 From previous section, we can clearly identify the limiting factors for energy efficient CO₂
37 conversion in the GA. Therefore, in this section, we will propose solutions on how to further improve
38 the performance of the GA for energy efficient CO₂ conversion. First we will discuss the role of the
39 vibrational levels in energy efficient CO₂ conversion. Subsequently, we will look in more detail at the
40 recombination of CO with O₂, which contributes mostly to the CO₂ formation at the conditions under
41 study. Finally, we will elaborate on some ways to increasing the velocity difference between the GA
42 and the gas flow, which can increase the fraction of CO₂ that can be processed by the arc, and hence
43 improve the conversion.

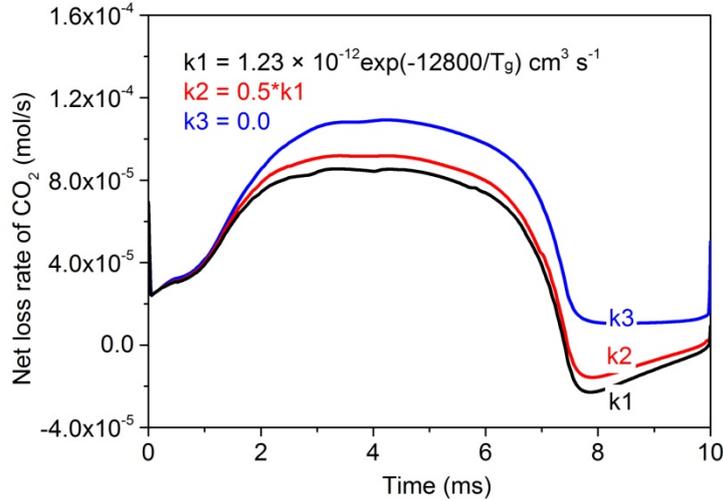
1 **4.5.1 Promoting the vibrational kinetics**

2 It is clear that non-equilibrium vibrational excitation of CO₂ promotes energy efficient
3 dissociation in the GA. This is also consistent with experimental investigations in literature [6]. Our
4 results indicate that the population of the symmetric mode levels and the lower asymmetric stretch
5 mode levels is much higher than that of the higher asymmetric mode levels. Therefore, these lower
6 vibrationally excited levels mostly account for the total CO₂ conversion, although there is still some
7 overpopulation for the higher levels. The reason why especially the lower vibrational levels
8 contribute to the CO₂ conversion is because the vibrational energy distribution function tends to
9 become more thermalized at high gas temperature [73]. Indeed, the energy exchange upon collision
10 between vibrational levels and ground state molecules, which depopulates the vibrational levels, i.e.,
11 so-called VT relaxation, increases with gas temperature. Therefore, we should look for ways of
12 inhibiting the VT relaxation process to increase the degree of overpopulation of the higher
13 asymmetric mode levels.

14 A recent kinetic modelling of microwave plasma based CO₂ conversion has shown that lower
15 pressures, lower gas temperature and higher power densities (at least for pressures below 300 mbar)
16 lead to more vibrational excitation, which is beneficial for the conversion [73]. However, our GA
17 operates at atmospheric pressure, which is more convenient for industrial applications, so the
18 solutions of reducing the gas pressure and increasing the power density (which only has beneficial
19 effect at a pressure below 300 mbar [73]) are not practical. Therefore, we believe that the gas
20 temperature should be reduced, to inhibit the VT relaxation, and thus to promote the role of the
21 higher vibrational levels, and hence the conversion and energy efficiency. In this respect, enhancing
22 the mixing between the GA and the cold gas can help to realize this goal, which was clearly indicated
23 by our previous modelling for a 1D gliding arc [41] and by experimental work [6]. Furthermore,
24 reducing the gas temperature will also result in a lowering of the recombination reactions, thus also
25 improving the overall CO₂ conversion (see next section). On the other hand, it will also lead to a drop
26 in the dissociation rate constants by neutral particle collisions, and this has a detrimental effect on
27 the conversion. Therefore, an optimized gas temperature should exist for GA based CO₂ conversion,
28 where the beneficial effect of a lower temperature, due to (i) a more pronounced non-equilibrium
29 population of the highly excited vibrational levels, and (ii) lower recombination rates of CO back into
30 CO₂, exceeds the detrimental effect by the lower dissociation rate constants of dissociation upon
31 collision with neutral particles. Finding out this optimal temperature is, however, not so
32 straightforward with our 2D model, as the latter self-consistently calculates the gas temperature and
33 it is not an input in the model. For this purpose, a 0D model, where the gas temperature can be
34 introduced as an input parameter, could be more suitable [40].

35 Besides, because electron impact vibrational excitation of CO₂ is mainly important for reduced
36 electric field values (i.e., ratio of electric field over gas density) below 80 Td [72] (where 1 Td = 10⁻²¹
37 V/m²), we should target to actively tune the reduced electric field to these values, by optimizing the
38 reactor electrical operating parameters. Finally, increasing the electron number density will also
39 promote the vibrational excitation and thus selectively deliver energy to this most energy efficient
40 CO₂ dissociation pathway. It has been reported in literature [66] that adding noble gases, such as
41 argon, to CO₂ would improve the CO₂ conversion and energy efficiency by increasing the electron
42 number density, because argon has a lower breakdown voltage than CO₂.

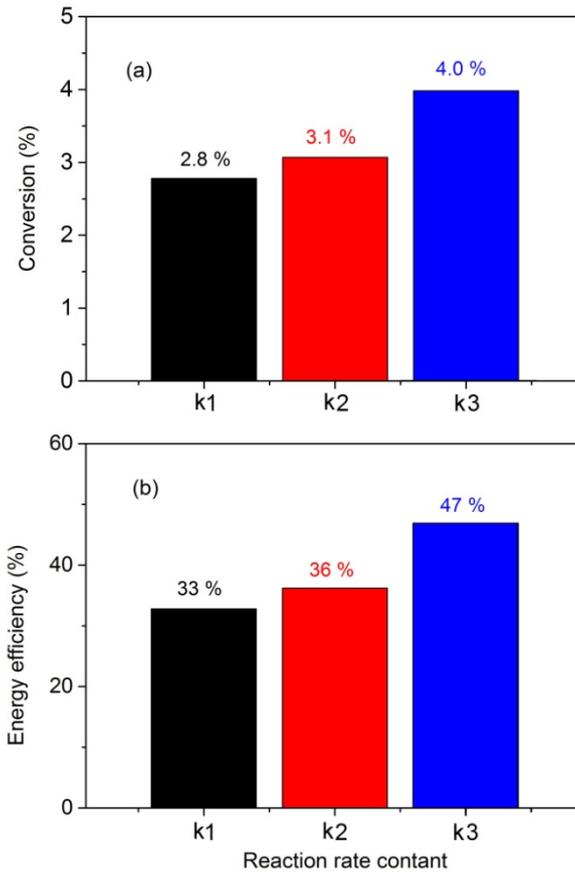
43 **4.5.2 Reducing the recombination of CO with O₂**



1
2 **Figure 7** Effect of different rate coefficients of the recombination reaction ($\text{CO} + \text{O}_2 \rightarrow \text{CO}_2 + \text{O}$) on the
3 calculated net loss rate of CO_2 , integrated over the entire reactor volume, at the same conditions as in
4 figure 4.

5 It is clear from section 4.4 that the recombination reaction (F1), i.e., $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2 + \text{O}$, is
6 mainly limiting the CO_2 conversion and energy efficiency. In our model, we adopted the rate
7 coefficient as proposed by Fridman [19]. However, to evaluate the effect of this recombination
8 reaction on the overall CO_2 conversion, we have performed some further simulations in which (i) we
9 reduced the rate coefficient of this reaction by 50%, and (ii) we completely removed this
10 recombination reaction from the model, as indicated in the legend of figure 7.

11 It is obvious from figure 7 that a lower rate coefficient of the recombination reaction yields a
12 higher net CO_2 loss rate. The CO concentration within the GA channel, and hence the influence of the
13 recombination reaction on the CO_2 formation, is minor till $t = 1.7$ ms. As a result, the different rate
14 coefficients have a negligible effect on the net loss rate of CO_2 up to 1.7 ms. Upon increasing CO
15 concentration, the different rate coefficients do cause some deviation in the calculated net loss rates
16 of CO_2 . After $t = 7.5$ ms, the formation rate of CO_2 is even larger than the loss rate for k_1 and k_2 ,
17 leading to a negative value of the net CO_2 splitting rate. Of course, integrated over the entire GA
18 cycle, the overall CO_2 loss (or conversion) rate is still positive, but it is greatly reduced due to this
19 important backward (recombination) reaction.



1
 2 **Figure 8** Effect of using different rate coefficients of the recombination reaction ($\text{CO} + \text{O}_2 \rightarrow \text{CO}_2 + \text{O}$)
 3 on the calculated CO_2 conversion (a) and energy efficiency (b), for the same conditions as in figure 7.
 4 See legend of figure 7 for the values of k_1 , k_2 and k_3 .

5 Figure 8 shows the conversion and energy efficiency, calculated with the original rate coefficient
 6 (k_1) [19], in comparison with the results obtained when this rate coefficient is divided by 2 (k_2), as
 7 well as when the recombination reaction is removed from the model (k_3). The conversion and energy
 8 efficiency increase only slightly when the recombination rate coefficient is divided by 2, while they
 9 rise from 2.8 % to 4.0 %, and from 33 % to 47 %, respectively, by removing the recombination
 10 reaction ($\text{CO} + \text{O}_2 \rightarrow \text{CO}_2 + \text{O}$) from the model. Although the conversion is still low, the energy
 11 efficiency rises significantly. This clearly indicates that reducing the recombination of CO with O_2 is
 12 quite promising to enhance the CO_2 conversion and (especially) the energy efficiency.

13 To achieve this objective, we suggest to apply possible scavengers, catalysts or separation
 14 membranes, in order to remove the O_2 molecules [31]. These are only suggestions, and they should
 15 of course be experimentally explored to evaluate the possibilities. On the other hand, the
 16 combination of a solid oxide electrolyser cell with a plasma set-up was already illustrated in [74] to be
 17 beneficial for the CO_2 conversion, and it works according to the same principle. In this way, the local
 18 concentration of O_2 molecules within the arc channel, and hence the net formation of CO_2 by the
 19 recombination reaction ($\text{CO} + \text{O}_2 \rightarrow \text{CO}_2 + \text{O}$), could be reduced, because there is not enough reactant
 20 (O_2) available for the backward reaction from CO into CO_2 (F1).

21 H_2 or CH_4 could act as possible scavengers for atomic oxygen, forming H_2O . This possibility was
 22 already illustrated to be beneficial for O trapping in literature, based on a combined plasma chemical

1 kinetics model and experiments for CO₂ conversion in another type of plasma [75]. The trapping of O
2 atoms might be able to promote the CO₂ conversion by (i) inhibiting the recombination reaction F2
3 [40], and (ii) by avoiding the formation of O₂, which will inhibit the recombination reaction F1.
4 Experiments in literature have indeed revealed that the addition of H₂ or CH₄ in a GA reactor can
5 improve the conversion of CO₂ [6],[54], but the enhanced conversion of CO₂ cannot be simply, or
6 entirely, attributed to the inhibited recombination reactions. This is because the H atoms or CH_x
7 radicals produced by H₂ or CH₄ dissociation can also contribute to CO₂ dissociation. Moreover, the
8 removal of O atoms will also inhibit the dominant mechanism of CO₂ splitting, i.e. the dissociation of
9 CO₂ upon collision with O atoms (L2v, L2g) and thus it might also exhibit a negative effect on further
10 improving the CO₂ conversion. Therefore, the reason why adding H₂ or CH₄ promotes the CO₂
11 conversion is not necessarily attributed to their scavenging role in consuming the O atoms. Indeed,
12 the direct involvement in CO₂ splitting by the reversed water gas shift reaction (CO₂ + H₂ → CO + H₂O)
13 has been verified to be a very important path for CO₂ splitting into CO when CH₄ [54] or H₂ [6] is added
14 into a CO₂ GA plasma. Moreover, the addition of H₂ or CH₄ can increase the electron density by
15 inhibiting electron attachment to O₂ (which is an electronegative gas), and this can also contribute to
16 a higher CO₂ conversion.

17 The idea of using a catalyst with a high surface interaction for O atoms to recombine into O₂
18 [76] or for O₂ adsorption is probably not very effective, because the O₂ molecules would be released
19 back to the plasma phase and again undergo recombination with CO. In contrast, a more advanced
20 catalytic process would be an alternative form of chemical looping, in which the O or O₂ is captured in
21 the plasma set-up and then used as oxidizing agent in a second set-up [77]-[78]. However, this is only
22 a concept, and has not been demonstrated yet for a GA reactor.

23 The third method, based on separation membrane technology, would transport the O₂
24 molecules (or O atoms) away from the reaction mixture. For example, by combination of a solid oxide
25 electrolyser cell with a plasma set-up, Tagawa et al. [74] and Mori et al. [79]-[80] have observed an
26 increasing CO₂ conversion by placing an O₂ trapping membrane into a CO₂/CH₄ or CO₂ discharge, in
27 order to separate O₂ from the reaction mixture.

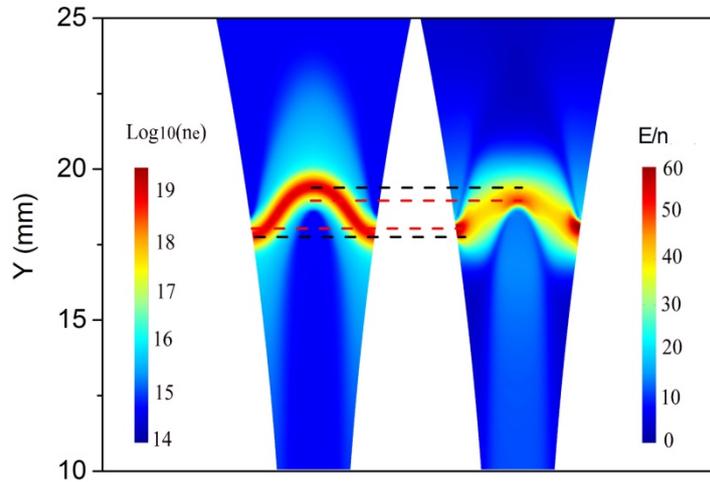
28 Besides the effect of possible scavengers, catalysts or membranes to remove the oxygen, as
29 mentioned above, we believe that the recombination of CO with O₂ could also be avoided or
30 minimized by providing effective quenching of the high temperature in the arc zone, due to mixing
31 with cold gas at very fast cooling rates. This could be especially beneficial in the relaxation stage of
32 the GA (around 8 ms) when the discharge current is low, and the CO₂ loss rate is minor, but the
33 recombination rate of CO with O₂ is still very large due to the very high gas temperature, leading to
34 net CO₂ formation. Indeed, an effective quenching of the residual plasma temperature can help to
35 decrease the recombination reaction rate and inhibit the CO₂ formation in this stage, leading to an
36 improved conversion and energy efficiency. We believe that such a quenching of the plasma
37 temperature could be realized by improving the reactor geometry and/or optimizing the flow
38 conditions, but further studies are needed to elaborate on these solutions.

39 **4.5.3 Increasing the CO₂ fraction to be treated by the arc due to a velocity difference** 40 **between GA and gas flow**

41 Besides promoting the vibrational kinetics and reducing the recombination reaction of CO into
42 CO₂, another way to improve the CO₂ conversion would be to enhance the CO₂ fraction to be treated

1 by the arc, by better mixing of the GA and the cold gas flow. This can be realized when there is a
2 velocity difference between the GA and the gas flow. Several experimental studies indeed have
3 shown that the arc gliding velocity can be slightly lower than the gas velocity [14], [33]. We present
4 here some simulation results, showing that there can indeed be a (small) difference between the arc
5 and gas flow velocity. We can distinguish two different ways to realize this.

6 (1) Smooth velocity difference due to the arc bending



7
8 **Figure 9** 2D distribution of the electron number density (left, in m^{-3}) and reduced electric field (right,
9 in Td) at a time instant of 2.5 ms for the same conditions as in figure 4. The black and red lines
10 indicate the position of the arc center and of the maximum reduced electric field, respectively,
11 showing that they are separated, leading to extra ionization downstream the arc in the centre, and
12 consequently to slowing down of the arc movement.

13 The first possible reason for a lower arc velocity vs gas flow velocity is related to the arc bending,
14 and thus the existence of zones with increased electric field outside the arc centre. The latter indeed
15 leads to a separation of the arc centre (with the maximum electron number density) and the position
16 with maximum reduced electric field, as presented in figure 9. This is caused by the fact that in the
17 symmetry plane, when the arc is highly bended, some parts of the arc in the downstream region of
18 the arc centre are positioned closer to each other. This increases the electric field strength in this
19 region and causes a gradual ionization of the gas in the downstream region. The latter will result in a
20 slightly lower arc velocity compared to the gas velocity. Likewise, near the walls (cathode and anode),
21 the maximum reduced electric field, and hence the gradual ionisation, appears in the upstream
22 region of the arc centre, which results in a slightly higher arc velocity than the gas velocity. Thus, the
23 GA moves a bit slower than the gas flow in the central part of the reactor and a bit faster in the
24 regions near the walls. At $t = 2.5$ ms, our calculation predict a GA velocity of 5.9 m/s in the centre,
25 compared to a gas flow velocity of 7.4 m/s. The ratio of gas velocity to arc velocity is thus 1.2, which
26 is in reasonable agreement with experiments [14], [33].

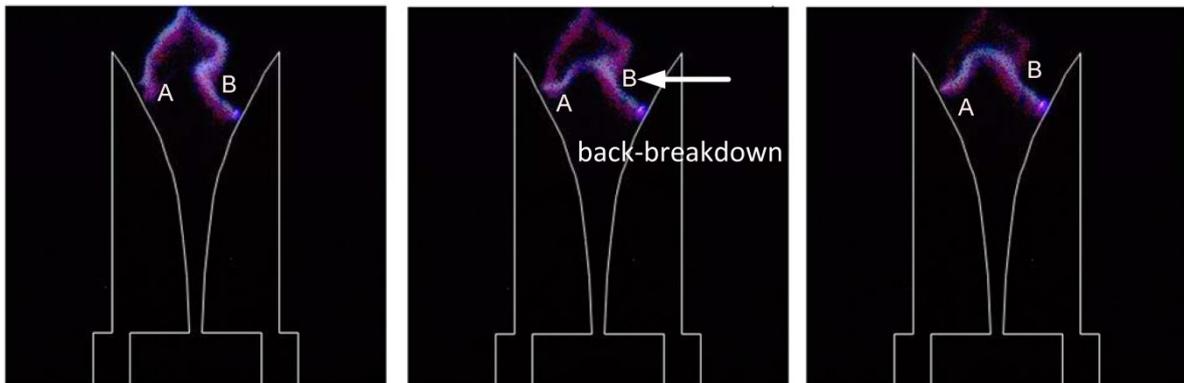
27 We have also performed calculations at higher gas flow velocity, and the results show that this
28 leads to an increased velocity difference between the arc and gas flow. For example, with the same
29 gas flow rate of 2.5 L/min, assuming the flow passing through a channel with depth of 1 mm, which is
30 only half of the value in our standard model, the gas flow velocity at the same time instant $t = 2.5$ ms
31 was calculated to be 8.3 m/s, with a GA velocity of 5.8 m/s, thus yielding a ratio of gas velocity vs arc

1 velocity of 1.4. This clearly shows that the velocity difference between GA and gas flow will be higher
2 for higher gas flow velocities, which is also reported in experiments [14], [33]. Correspondingly, our
3 calculated conversion increases from 2.78 % to 4.4 %, although the energy efficiency only increases
4 from 32.8 % to 34 %. Although this is an artificial method, we can show in this way that the
5 treatment capacity can be enlarged by increasing the local gas velocity and hence the relative
6 velocity between gas flow and GA. Increasing the local gas velocity can be realized by modifying the
7 reactor setup and hence the flow configuration at a fixed gas flow rate, for example by shortening
8 the narrowest gap separation of both electrodes [40] or by reducing the distance between the nozzle
9 exit and the reactor [81] or by decreasing the nozzle internal diameter [82]. Indeed, following such
10 methods, increased conversions were reached experimentally [40], [81] and [82]. However, we should
11 also mention that simply adjusting these parameters is not a proper way to enhance the treatment
12 capacity of the GA reactor, because it might give rise to an extreme increase in the gas velocity,
13 which may greatly reduce the effective residence time of CO₂ in the GA volume. This is of course
14 detrimental for the CO₂ conversion. Moreover, the high gas velocity will bring a strong cooling effect
15 and hence a lower gas temperature; the latter can be beneficial (to promote the vibrational kinetics
16 and/or reduce the recombination reactions), but it may also be detrimental (due to the reduced
17 dissociation reaction rate constant), as we discussed in section 4.5.1 above. Therefore, the above
18 mentioned operating parameters should be optimized in a suitable range, to guarantee an
19 improvement in conversion and energy efficiency [83].

20 (2) Sudden velocity difference due to back-breakdown events

21 Besides the smooth reduction in GA velocity explained above, another reason for the lower arc
22 velocity vs gas flow velocity is related to the instabilities of the arc and to secondary breakdowns, also
23 called back-breakdown, causing a reduction in arc length [84].

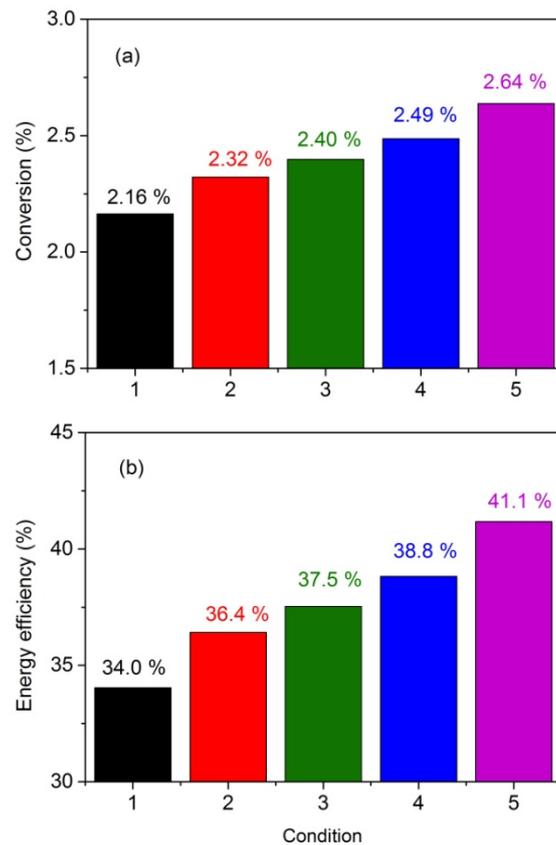
24 The back-breakdown phenomena, which result in a fast shortening of the arc as a result of
25 breakdown between different parts of the arc (instead of between the electrodes) often take place in
26 a GA, especially at higher gas flow rates, as also mentioned in section 4.1. These shortcuts effectively
27 appear as a lag of the arc velocity compared to the gas flow and could be an efficient mechanism for
28 the treatment of a larger gas fraction. As explained at the end of section 4.1, this effect is not taken
29 into account in previous sections, because at the gas flow rate of 2.5 L/min, our high speed camera
30 did not record any back-breakdown events.



31
32 **Figure 10** Back-breakdown event recorded by the high speed camera at a flow rate of 5 L/min (5000
33 frames/s, exposure time of 50 μ s, electrode throat of 2.0 mm)

1 Figure 10 illustrates a back-breakdown event, recorded by the high speed camera at a flow rate
 2 of 5 L/min. Indeed, at a high gas flow rate (above 2.5 L/min), the GA discharge is unstable and it has a
 3 rather irregular shape. When some parts of the GA (see points A and B in figure 10) get closer to each
 4 other, the electric field there increases. Once the potential difference between these two parts, and
 5 hence the local electric field, exceeds the critical breakdown electric field [85], a new discharge
 6 channel is established (see middle panel) and the old discharge channel disappears very fast. This
 7 causes a drop in the GA velocity as compared to the gas flow velocity.

8 Although several experiments [36],[84] have been performed to study the back-breakdown
 9 events, it is not straightforward to establish a self-consistent back-breakdown model, since this
 10 behaviour is mostly stochastic by nature and the arc instabilities are not well defined. To investigate
 11 here the influence of the back-breakdown events on the CO₂ conversion, we have initiated this
 12 process by establishing an artificial plasma channel, which is triggered on a regular or irregular basis
 13 with respect to the arc path or time, i.e., after every certain distance or period. Details on how the
 14 back-breakdown model is established can be found in [38], as well as in the supporting information
 15 of our paper.



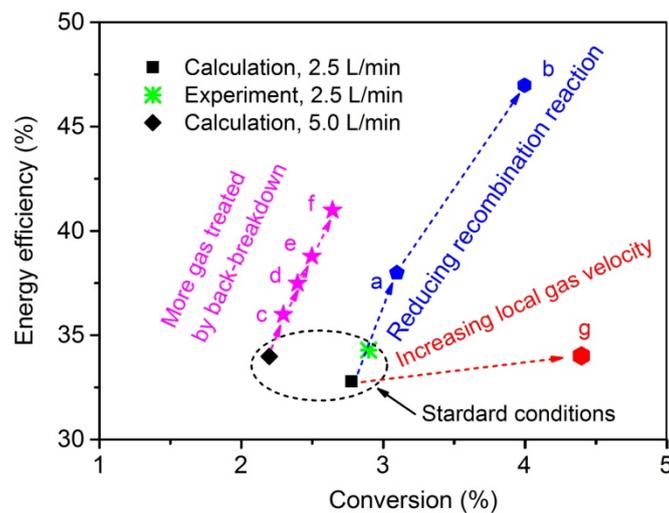
17 **Figure 11** Effect of the back-breakdown events on the CO₂ conversion and energy efficiency at a gas
 18 flow rate of 5.0 L/min, for different cases, i.e., without back breakdown (1); one back-breakdown at 5
 19 ms (2); a two back-breakdowns, at 5 ms and 6 ms (3); three back-breakdowns, at 5 ms, 6 ms and 7 ms
 20 (4); and five back-breakdowns, at 5 ms, 5.5 ms, 6 ms, 6.5 ms and 7 ms (5).
 21

22 Figure 11 illustrates the effect of the back-breakdown events on the calculated conversion and
 23 energy efficiency. The power needed to initiate the back-breakdown events is included in the
 24 determination of the total plasma power and hence in the SEI value in Eq. (3), as well as the

1 calculation of the energy efficiency in Eq. (4) (see section 2). It is clear that the back-breakdown
 2 events yield an improved CO₂ conversion and energy efficiency, compared with the case without
 3 back-breakdown, because a larger fraction of CO₂ is treated by the newly established discharge
 4 channel. This also explains why a larger number of back-breakdown events can enhance the CO₂
 5 conversion and energy efficiency (see cases 2, 3, 4 and 5). Moreover, more back-breakdown events
 6 also result in a lower overall gas temperature, as is clear from figure S7 of the supporting information,
 7 because the heat is now spread over a larger domain and not only within the initial arc channel. This
 8 lower gas temperature can have beneficial or detrimental effects on the overall CO₂ conversion, as
 9 explained above.

10 As discussed above, the occurrence of the back-breakdown events is closely linked with two
 11 factors, i.e. the arc instabilities and a sufficiently high arc voltage drop. The former leads to a rather
 12 irregular arc shape and a non-stable discharge, increasing the probability of a closer interaction
 13 between two separated parts of the GA. The latter can ensure a high enough electric field between
 14 the two separated parts of the arc, to ignite a new discharge channel. In order to satisfy these two
 15 essential requirements, besides increasing the gas flow rate, the gas flow velocity must also be
 16 increased by modifying the reactor setup and hence the flow configuration under a fixed gas flow
 17 rate, as discussed above.

18 4.5.4 Summary of the proposed improvements



19
 20 **Figure 12** Energy efficiency vs CO₂ conversion in our GA reactor, as obtained by our experiments and
 21 calculated by our model, for the standard conditions (indicated with the oval), as well as several
 22 improvements as predicted by the model, by either (i) reducing the recombination rate coefficient
 23 from k_1 to k_2 (a) and k_3 (b) (cf. figure 7), or (ii) enhancing the treated CO₂ fraction, by increasing the
 24 number of back-breakdown events, from 1 (c) to 2 (d) to 3 (e) to 5 (f), applicable at a higher gas flow
 25 rate (5 L/min), or (iii) by increasing the local gas velocity at the same gas flow rate, due to reactor
 26 inlet modifications, leading to a higher velocity ratio between gas flow and GA (g).

27 Finally, in figure 12, we schematically summarize the improvement in the CO₂ conversion and
 28 energy efficiency, as proposed and predicted by our model. The CO₂ conversion and energy efficiency
 29 are about 2.78 % and 32.8 % (calculated) or 2.90 % and 34.3 % (measured) at the standard conditions
 30 investigated, i.e., a gas flow rate of 2.5 L/min and a plasma power of 40 W, corresponding to a SEI of
 31 0.25 eV/molecule. However, these values can be improved according to the model predictions, up to

1 a conversion of nearly 4 % and a corresponding energy efficiency of 47 % (see point b) by inhibiting
2 the recombination reaction of CO with O₂. Furthermore, if the gas fraction that can pass through the
3 arc zone could be enhanced, for instance by modifying the reactor setup and hence the flow
4 configuration to realize a higher relative velocity between arc and gas flow, the conversion and
5 energy efficiency are predicted to increase to 4.4 % and 34 %, respectively (see point g). Finally, the
6 occurrence of back-breakdown events, which induce an abrupt difference in gas flow velocity and GA
7 velocity, in case of a gas flow rate of 5 L/min (where the back-breakdown events indeed can take
8 place), can also help to increase the conversion, although the effect seems to be rather limited, with
9 a maximum conversion up to 2.6 %, while the energy efficiency would increase up to 41 % (see point
10 f).

11 The proposed solutions yield some improvement in conversion and energy efficiency, but these
12 model predictions still need to be verified by experiments. We hope that our model predictions will
13 inspire experimental researchers to try out these modifications. Furthermore, the improvements are
14 probably still too limited for industrial application of the GA for CO₂ conversion. Indeed, although the
15 energy efficiency is quite good, the conversion is still very limited. Hence, more drastic modifications
16 would be needed, e.g., in the gas flow pattern or the source design, to significantly increase the
17 fraction of gas that can pass through the arc. One possible suggestion would be the reverse vortex
18 flow gliding arc, which is based on cylindrical electrodes, and which allows a larger fraction of the gas
19 to pass through the arc, yielding higher CO₂ conversions, as demonstrated by [6], [50] and **Error!**
20 **Reference source not found..**

21 **2 Conclusions**

22 In this work we studied the CO₂ conversion in a GA plasma, by means of a combined
23 experimental and 2D modelling approach. We compared our measured and calculated CO₂
24 conversion and corresponding energy efficiency, as well as the electron number density in the arc,
25 and obtained reasonable agreement. This indicates that our model can provide a realistic picture of
26 the plasma chemistry and can be used to elucidate the underlying mechanisms and the dominant
27 reaction pathways for the GA based CO₂ conversion.

28 We presented the typical arc plasma characteristics, such as the electron number density,
29 electron temperature and gas temperature, as well the CO molar fraction, for one entire arc gliding
30 cycle, as calculated by our model. These results clearly show that the GA plasma has a strong non-
31 equilibrium character, because the electron temperature is much higher than the gas temperature,
32 and the highly energetic electrons can induce several different chemical reactions. This explains the
33 better performance of the GA for CO₂ conversion, yielding a much higher energy efficiency for a fixed
34 value of the conversion, than pure thermal conversion, for which the energy is distributed over all
35 degrees of freedom, including those not effective for the CO₂ conversion.

36 We also performed a chemical kinetics analysis of the modelling results, which enables us to
37 identify the important species and reactions playing a role in the CO₂ splitting, i.e., the main
38 production and loss pathways of CO₂. This allows us to gain sufficient insight into the entire process,
39 and to identify the limiting factors for CO₂ conversion, and thus to propose solutions for improving
40 the CO₂ conversion. Our model predicts that the most important process for CO₂ conversion is the
41 dissociation of vibrationally excited states of CO₂ upon collision with O atoms, indicating that the CO₂
42 vibrational levels significantly contribute to the CO₂ dissociation. This can explain the good energy
43 efficiency of CO₂ conversion in a GA plasma, as compared to some other plasma types.

1 We believe that, when it is possible to actively tune the reduced electric field (i.e., E/n ratio) in
2 the plasma, by optimizing the reactor electrical operating parameters, or when we can increase the
3 electron number density, as well as inhibit the VT relaxation processes by decreasing the gas
4 temperature, we should be able to further promote the vibrational excitation and selectively deliver
5 energy to the CO₂ dissociation via this energy efficient pathway. This should lead to some further
6 improvement in the energy efficiency of CO₂ conversion in the GA.

7 Furthermore, our calculation shows that the reverse reactions, especially the recombination of
8 CO with O₂ molecules (and to a lower extent with O atoms), have a non-negligible rate, compared to
9 the CO₂ loss rate. Therefore, these reactions have a detrimental effect on the overall CO₂ conversion.
10 Thus, in order to further improve the CO₂ conversion, the reversion reactions should be inhibited or
11 at least reduced. We clearly demonstrate this by running the model with different reaction rate
12 coefficients for recombination, and when this recombination reaction is entirely removed, the
13 calculated CO₂ conversion and energy efficiency rise from 2.8 % and 33 %, to 4.0 % and 47 %,
14 respectively.

15 Finally, our simulation shows that the molar fraction of CO₂ within the arc center is very low,
16 indicating that the local CO₂ conversion is nearly complete, but because the fraction of treated CO₂
17 within the arc is very limited, the overall CO₂ conversion is also limited. Therefore, we should look for
18 ways to increase the CO₂ fraction to be treated by the arc, in order to further improve the GA based
19 CO₂ conversion. Increasing this treated gas fraction can be realized when there is a velocity
20 difference between the GA and the gas flow, so that new fractions of the CO₂ gas can pass through
21 the arc, while the converted fraction (i.e., CO, O and O₂) will leave the active arc region, before it can
22 recombine back into CO₂. We therefore discuss possible ways of increasing the relative velocity
23 between GA and gas flow. The first way to realize this is by increasing the local gas velocity without
24 changing the gas flow rate, for instance by modifying the reactor setup and hence the flow
25 configuration. Indeed, at a high gas velocity, there is a larger difference between GA and gas flow
26 velocity due to some ionization downstream the arc channel, slowing down the arc movement.
27 Additionally, the occurrence of back-breakdown events, creating new conducting arc channels, will
28 also cause a difference between GA and gas flow velocity, so we also investigated the effect of these
29 back-breakdown events on the calculated CO₂ conversion and energy efficiency. Our calculations
30 clearly indicate that the back-breakdown events, which generally take place at a high gas flow rate,
31 can help to further increase the CO₂ conversion and energy efficiency.

32 This study is of great interest for GA based CO₂ conversion, as we were able to elucidate the
33 main underlying mechanisms and chemical reactions of the conversion process by means of a model
34 that was validated by experiments. In general, we illustrated that GA based CO₂ conversion is quite
35 promising, when compared with the classical thermal CO₂ conversion process, as well as with other
36 plasma types. This is attributed to its non-equilibrium character, promoting the vibrational kinetics.
37 However, we believe there is still room for improvement. Indeed, we could identify the limiting
38 factors of the CO₂ conversion in the GA, and thus propose solutions on how to further improve the
39 performance.

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