Oxidative Desulfurization of Model Diesel Fuel Catalyzed by Carbon-Supported Heteropoly acids

Reem Ghubayra, Charlotte Nuttall, Sophie Hodgkiss, Michael Craven, Elena F. Kozhevnikova, Ivan V. Kozhevnikov*[[1]](#footnote-1)*

*Department of Chemistry, University of Liverpool, Liverpool, L69 7ZD, United Kingdom*

**Abstract**

Keggin-type heteropoly acids supported on activated carbon (HPA/C) are active catalysts for oxidative desulfurization (ODS) of diesel fuel under mild conditions in a biphasic system composed of a benzothiophene-containing model diesel fuel (heptane) and aqueous 30% H2O2. The catalytic activity of HPA/C was found to decrease in the order of HPA: H3PMo12O40 > H3PW12O40 > H4SiW12O40. The most active catalyst, H3PMo12O40/C, exhibited 100% removal of benzothiophenes from model diesel fuel at 60 oC, and could be recovered and reused without loss of activity. This catalyst outperforms other recently reported heterogeneous catalysts for ODS in similar systems. Kinetic and DRIFTS studies provide new insights into the mechanism of ODS reaction on carbon-supported HPAs.

**Keywords:** oxidative desulfurization; heteropoly acid; carbon support; heterogeneous catalysis

**1. Introduction**

The sulfur content of diesel fuels used in transportation vehicles is under increasingly strict environmental regulations. It has decreased from 2000 to 10 ppm over the last 20 years and further reduction in sulfur content is desired [1,2]. Hydrodesulfurization (HDS), the most widely used technology for removing sulfur from diesel fuels, is operated at high temperature (300-400 oC) and pressure (30-130 atm) using alumina-supported Co-Mo or Ni-Mo catalysts [1,2]. The main drawbacks of HDS are severe operating conditions and low desulfurization efficiency in the case of refractory benzothiophenes. Alternative desulfurization methods have been investigated in recent years including, among others, oxidative desulfurization (ODS) [3,4], extraction [5], adsorption [6] and bio-desulfurization [7]. ODS appears to be the most promising method for deep desulfurization of diesel fuel. Typically, it involves liquid-phase biphasic oxidation of organosulfur compounds with H2O2 at low temperatures (50-70 oC) and atmospheric pressure to yield sulfoxides and sulfones, which can be separated from the fuel by precipitation, extraction or adsorption [3,4]. This method is highly efficient for removing refractory aromatic sulfur compounds such as thiols and benzothiophenes, which are difficult to remove by HDS [2].

Among many reported ODS catalysts, polyoxometalates (POMs), in particular Keggin-type POMs, exhibit a remarkable activity [8-13]. These compounds comprise polyanions, XM12O40*m-,* composed of oxygen-sharing MO6 octahedra (M = MoVI, WVI, VV, etc.) encapsulating a central tetrahedron XO4*n-* (X = PV, SiIV, etc.) [13]. In solution, in the presence of hydrogen peroxide, these POMs degrade to form active peroxo-polyoxometalate species (peroxo-POM) [12,14,15], e.g., the Venturello peroxo complex, {PO4[WO(O2)2]4}3- [16]. Peroxo-POMs are highly active catalysts for various biphasic oxidations with hydrogen peroxide [17-20]. In these reactions, a phase transfer agent (PTA) is required to transfer the peroxo-POM species from the aqueous phase containing H2O2 to the substrate-containing organic or fuel phase. Most frequently, quaternary ammonium cations are used as the PTAs in such systems, including amino-modified high molecular weight alkene oligomers [20]. Alkylaminophosphazenes have also been reported as promising phase transfer agents in POM-catalyzed biphasic oxidations with H2O2, including the oxidative desulfurization of benzothiophenes [12,21].

A major drawback of homogeneous POM catalysts in the PTA-assisted biphasic oxidative desulfurization is the difficulty of separating these catalysts from the fuel phase after desulfurization because the POM-PTA aggregates are highly soluble in the fuel phase. This would lead to contamination of fuel with POM. In this regard, heterogeneous ODS catalysts have the important advantage of easy catalyst separation from the fuel after reaction [22-28]. However, possible leaching of POM from the solid catalyst during the ODS reaction and post reaction catalyst treatment is a serious problem of the heterogeneous ODS systems. This may be addressed by incorporating POM within a metal-organic framework [26] or by chemical immobilization of POM onto an appropriately functionalized support [27]. A more practical solution to overcome this problem could be the use of porous supports that can strongly adsorb the active POM species.

This study aims at exploring activated carbon as a support for the heterogeneous ODS reaction catalyzed by POMs. Previously, it has been reported that activated carbon can strongly adsorb Keggin heteropoly acids (HPA), such as H3PW12O40 and H4SiW12O40, irreversibly retaining 7 - 14 wt% of HPA in polar media such as water and methanol [29-31]. As a result, the corresponding HPA/C catalysts exhibit high stability towards HPA leaching in such media [30-32]. Strong interaction of HPA with carbon support is evidenced by a large line broadening in 31P MAS NMR spectra of H3PW12O40 supported on activated carbon [33]. From IR and 31P MAS NMR, heteropoly acids H3PW12O40 and H4SiW12O40 supported on a chemically (H3PO4) activated carbon retain the Keggin structure at an HPA loading >5 wt% [33]. As evidenced by XRD, heteropoly acids form finely dispersed species on the carbon surface; no HPA crystal phase is developed even at a HPA loading as high as 45 wt% [33]. Microcalorimetry of ammonia adsorption shows that the acid strength of H3PW12O40 is greatly reduced when loading on activated carbon [34,35]. This suggests HPA bonding through protonation of the carbon framework. Scanning tunnelling microscopy and tunnelling spectroscopy studies of HPA supported on graphite have been reported [36]. Carbon-supported POMs are well-documented catalysts for liquid-phase oxidation [37-39]. Recently, a polyoxometalate [PSPy]3PMo12O40 supported on graphite, where PSPy is N-(3-sulfonatepropyl)-pyridinium ion, has been reported as an efficient ODS catalyst [28].

This work investigates heterogeneous catalysis by Keggin-type HPAs (H3PMo12O40, H3PW12O40 and H4SiW12O40) supported on activated carbon (HPA/C) for the biphasic oxidative desulfurization of model diesel fuel (heptane) by aqueous 30% H2O2. We report a highly efficient catalyst, H3PMo12O40/C, for the oxidation of benzothiophenes (benzothiophene (BT), dibenzothiophene (DBT) and 4,6-dimethyldibenzothiophene (DMDBT)) to the corresponding sulfones by H2O2. This catalyst shows a higher catalytic activity for the oxidation of benzothiophenes than other recently reported heterogeneous ODS catalysts in similar systems [22-27]. Strong adsorption of H3PMo12O40 onto the carbon support stabilizes the HPA structure and prevents the HPA from leaching. In addition, our kinetic and infrared spectroscopic studies provide new insights into the mechanism of ODS reaction on carbon-supported HPAs.

**2. Experimental**

*2.1. Chemicals*

Benzothiophene (BT, 99%), dibenzothiophene (DBT, 99%), 4,6-dimethyldibenzothiophene (DMBDT, 97%), heptane (99%), dodecane (99%), 30% H2O2, heteropoly acid hydrates H3PW12O40 (99%), H3PMo12O40 (99.9%) and H4SiWO40 (99.9%) containing 20–28 H2O molecules per Keggin unit were all purchased from Sigma–Aldrich. Darco KB-B activated carbon wet powder (150 μm particle size) was also from Sigma–Aldrich. It contained <100 ppm of Fe and 0.5% of extractable phosphates. From our thermogravimetric analysis (TGA), the activated carbon contained 17% of water (Supplementary Information, Fig. S1). Its elemental composition (%): C, 72.0; H, 3.0; N, 0.0; S, 0.0; O, 25.0 (by difference). Infrared spectrum of the activated carbon is shown in Fig. S2.

*2.2. Catalyst preparation*

The carbon-supported HPA catalysts (HPA/C) were prepared by impregnating Keggin-type heteropoly acids H3PMo12O40 (HPMo), H3PW12O40 (HPW) and H4SiW12O40 (HSiW) onto Darco KB-B activated carbon from aqueous solutions. The amount of crystallisation water in the commercial HPA hydrates was determined by TGA and taken into account in catalyst preparation. The HPA hydrates (0.15 g of anhydrous HPA) were placed in a 200-mL glass beaker and dissolved in a minimum amount of distilled water. Darco KB-B activated carbon (0.85 g) was added to the beaker and the mixture was magnetically stirred for 3 hours at 40 oC. The beaker was then placed in an oven and the content was dried at 100 oC overnight to afford the catalysts as black powders. The HPA loading in the catalysts was in the range of 11-13 wt% as determined from Mo and W analysis by ICP-OES (inductively coupled plasma - optical emission spectrometry) after digesting catalyst samples in aqueous 15% KOH. From TGA, the catalysts thus prepared contained 8% of water (Fig. S3). The HPA/C catalysts were kept in a desiccator over calcined silica gel. The information about HPA/C catalysts is given in Table 1.

*2.3. Techniques*

BET (Brunauer-Emmett-Teller) analysis of catalyst samples was conducted on a Micrometrics ASAP 2010 instrument by measuring N2 physisorption at -196 oC. The samples were pre-treated at 220 oC under vacuum (1 Pa). DRIFT (diffuse reflectance infrared Fourier transform) spectra were recorded on a Nicolet Nexus FTIR spectrometer using powdered catalyst mixtures with KBr. The spectra were recorded at room temperature in the absorbance mode by averaging 254 scans in the range of 4000-500 cm-1 with a resolution of 4 cm-1. DRIFT spectra of pure HPAs were measured against a KBr background and the spectra of HPA/C samples against a mixture of KBr and Darco KB-B activated carbon as the background. OMNIC software was used for data processing. Thermogravimetric analysis (TGA) was carried out on a Perkin Elmer TGA-7 instrument. HPA loading in the catalysts was determined on a Spectro Ciros ICP-OES analyzer; the HPA/C samples were digested by boiling in aqueous 15% KOH.

*2.4. Reaction procedure*

The oxidation of benzothiophenes by 30% hydrogen peroxide in the presence of HPA/C catalysts (≤1 wt% per total reaction mixture) was carried in a two-phase system, heptane-H2O, with a water content of ~2 vol.% added with aqueous 30% H2O2. To a 50 mL jacketed glass reaction vessel equipped with a heat circulator, a magnetic stirrer and a reflux condenser were added, unless stated otherwise: HPA/C catalyst (0.0041 mmol of HPA), heptane (5.0 mL) and aqueous 30% H2O2 (1.5 mmol, the concentration of H2O2 was standardised by titration with KMnO4). The reaction mixture was stirred at 60 oC for 5 min to activate the catalyst. The stirrer was stopped and 5.0 mL of heptane stock solution containing BT, DBT or DMDBT (0.50 mmol) and dodecane (0.40 mmol, GC internal standard) was added to the reaction mixture. The stirrer was then switched on to start the reaction. Catalyst pre-activation at the reaction temperature (60 oC/5 min) prior to adding the substrate allowed to exclude any induction period that otherwise could have taken place [27]. Typically, the reaction was carried out at 60 oC and 1500 rpm stirring speed unless stated otherwise. Reaction rate did not depend on the stirring speed in the range of 500-1500 rpm (Fig. S4), which indicates no external diffusion limitation. The conversion of benzothiophenes was monitored by submitting aliquots of the organic phase (0.1 mL, centrifuged to separate the catalyst) for analysis by gas chromatography (GC) using the internal standard method (a Varian Chrompack CP-3380 gas chromatograph equipped with a flame ionization detector and a 25 m × 0.32 mm × 0.5 µm BP1 capillary column). The mean absolute percentage error in substrate conversion was ≤5%. The reaction products, benzothiophene sulfones, were not soluble in heptane and precipitated from the reaction mixture. To determine the sulfone yield, after the ODS reaction had reached completion, the reaction mixture was diluted with 1,2-dichloroethane, stirred at 60 oC for 5 min to dissolve the sulfones and analyzed using GC.

For catalyst reuse in DBT oxidation, the initial reaction was run using the above procedure for 1 h. After the reaction had reached completion, the catalyst was separated and washed with toluene and then with heptane using a centrifuge. Afterwards, the catalyst was returned to the reactor along with the required amounts of all other reaction components (30% H2O2, dodecane, DBT and heptane) for the next run.

**3. Results and discussion**

*3.1. Catalyst characterization*

The carbon-supported HPA catalysts were prepared using a simple wet impregnation procedure with HPA loading in the range of 11-13 wt% (Table 1). The HPA loading did not change after stirring the catalysts (0.5 g) in water (100 mL) at 60 oC for 1 h. This confirms the strong adsorption of HPAs onto the carbon surface, in agreement with previous reports [29-31], and ensures stability of the HPA/C catalysts toward HPA leaching during the ODS process.

The loading of HPAs onto activated carbon had a very little effect on the microporous texture of the support, as evident from the nitrogen adsorption/desorption isotherms shown in Fig. 1 for HPMo/C and Fig. S5-S8 for HPW/C and HSiW/C. It increased the surface area only slightly (3-10%), with the pore volume and pore diameter remaining practically unchanged (Table 1). This indicates that the HPA was mostly present on the outer surface rather than occupying the pores of carbon support.

As evidenced by DRIFT spectroscopy, despite strong HPA-support interaction, the HPAs retained the Keggin structure upon their loading onto the carbon support. For the HPMo catalyst, this is clearly seen in Fig. 2, which shows the DRIFT spectra for the fresh HPMo/C catalyst (spectrum 2) and bulk crystalline HPMo (spectrum 1) as a reference; the latter shows the four characteristic bands at 1062 cm-1 (P-O), 960 cm-1 (Mo=O), 884 cm‑1 (Mo-O-Mo corner-sharing) and 811 cm-1 (Mo-O-Mo edge-sharing), in agreement with the literature [40]. Close resemblance between the two spectra confirms the integrity of the Keggin structure in the carbon-supported HPMo. The same conclusion also applies to HPW/C and HSiW/C catalysts (their DRIFT spectra are presented in Fig. S9 and S10, respectively), in agreement with previous report [33].

It should be noted that there is a small (10-20 cm-1) blue shift of the edge-sharing M-O-M band in HPA/C catalysts in comparison to the corresponding bulk HPAs (Fig. 2, Fig. S9 and Fig. S10). This may be explained by protonation of the edge-sharing bridging oxygens in supported HPAs. It is well known that protons in the HPA crystal lattice link the neighboring polyanions through the oxygen atoms of the terminal M=O groups, as evidenced by single-crystal X-ray and neutron diffraction data for HPW hexahydrate [41]. On the carbon surface, however, no HPA crystal phase has been found even at HPA loadings as high as 45 wt% [33]; in this case, the HPA would rather exist as a molecular dispersion. In isolated HPA molecules, in contrast to the crystalline HPA, protons are likely to be localized on the edge-sharing bridging oxygens that have a larger negative charge [13,42].

*3.3. Oxidation of benzothiophenes by H2O2 catalyzed by HPA/C*

The HPA/C catalysts comprising Keggin-type HPAs (HPMo, HPW and HSiW) supported onto activated carbon were used for ODS reaction in a biphasic heptane-H2O (98:2 v/v) system, with heptane as a model diesel fuel and benzothiophene (BT, DBT or DMDBT, ~1 wt%) as an organosulfur compound. 30% H2O2 was used as an oxidant; unless otherwise stated, it was used in 50% excess to the stoichiometric amount to ensure 100% conversion of benzothiophenes. Most of the work was carried out with DBT because the oxidation of DBT is typically employed as a model reaction for testing desulfurization catalysts, hence making it easier to compare catalyst performance of different catalyst systems in the literature.

Representative results are given in Table 2. The catalytic activity of HPAs in the oxidation of DBT was found to decrease in the order HPMo > HPW > HSiW (entries 1-3 in Table 2, Fig. 3). The same activity trend has been found previously in biphasic systems with homogeneous [12] and heterogeneous [27] POM catalysts modified with alkylaminocyclophosphazenes. Under our conditions, with H2O2 added in 50% excess, the oxidation of DBT gave DBT sulfone as the final product in a 100±13% yield (determined by GC analysis), with only traces of DBT sulfoxide found. Therefore, the oxidation of DBT can be represented by Scheme 1.



**Scheme 1.** Oxidation of DBT to sulfone by H2O2.

Activated carbon is well known to adsorb benzothiophenes from hydrocarbon media, and this has been used for desulfurization of diesel fuel [6]. In our system, however, this played only a minor role as only 5% of DBT was adsorbed on the HPMo/C catalyst in the absence of H2O2 (Table 2, entry 5). Also, activated carbon itself has been reported to catalyze the oxidation of DBT by H2O2 [43]. This was found to take place in our system as well, contributing up to 30% of DBT conversion (Table 5, entry 6). No reaction between DBT and H2O2 was observed in a blank test at 60 oC in the absence of catalyst and neat carbon support.

The reactivity of benzothiophenes was found to decrease in the order DBT > DMDBT > BT (Table 2, entries 1, 7 and 8; Fig. 4). The same trend has also been observed in homogeneous and other heterogeneous ODS systems [3,4,8,9,12,28,44,45]. It can be attributed to the electron density on the S atom and steric effects of the methyl groups in DMDBT [9]. The electron density on the S atom increases in the order BT < DBT ≈ DMDBT, which can explain the lower reactivity of BT as compared to DBT and DMDBT. On the other hand, the S atom in DMDBT is sterically hindered by the two neighboring methyl groups rendering it less reactive compared to DBT [9].

The efficiency of ODS depended on the relative amount of H2O2 in the reaction mixture, increasing with an increase in the H2O2/DBT molar ratio. Increasing the amount of H2O2, however, led to a decrease of the reaction selectivity based on H2O2 (H2O2 efficiency) due to non-productive decomposition of H2O2 to form O2 and H2O. For the HPMo/C catalyst, an optimum H2O2/DBT molar ratio was found to be 3.0 (50% H2O2 excess), which gave a 100% DBT conversion at 60 oC in 0.5 h with an 80% H2O2 efficiency as determined by post reaction titration with KMnO4 (Table 2). At stoichiometric ratio H2O2/DBT = 2.0 and otherwise the same conditions, the reaction reached a maximum 98% conversion in 1.5 h, thus providing 98% H2O2 efficiency (Table 2, entry 4). The decrease in H2O2 efficiency with increasing the concentration of H2O2 could be the result of zero order of the ODS reaction in H2O2 (see below) and a positive order of the non-productive H2O2 decomposition in H2O2. The H2O2 efficiency was also found to decrease with increasing the reaction temperature above 60 oC due to a faster decomposition of H2O2.

Notably, the HPMo/C catalyst was found to be more active than its recently reported homogeneous analogue with hexylaminocyclophosphazene (HexPN) as a phase transfer agent [12]. The HPMo/C catalyst gave 100% DBT conversion in 0.5 h at 60 oC (Table 2), whereas PMo-HexPN gave only 79% conversion under the same reaction conditions and with the same HPMo loading [12]. This may be explained by adsorption of DBT onto the surface of carbon support, which could increase the local concentration of DBT in the vicinity of active catalyst sites.

The HPMo/C catalyst showed higher activity than other heterogeneous catalysts that have been reported recently for the oxidation of DBT with hydrogen peroxide in similar systems [24-28]. Unfortunately, the data on turnover reaction rates are limited in the literature. Therefore, the results from different reports are compared on the basis of DBT conversion focusing on parameters such as H2O2/DBT molar ratio, reaction temperature and reaction time (Table 3). Ti(IV) grafted onto silica gives 99% DBT conversion to sulfone in isooctane with 10-60% H2O2 at 60 oC in 8 h reaction time [24]. With hybrid catalysts comprising Zr(IV) and Hf(IV) oxoclusters in poly(methylmethacrylate) matrix, 84% DBT conversion to sulfone with 94% selectivity has been obtained in octane with 30% H2O2 at 65 oC in 24 h [25]. A polyoxometalate-MOF composite comprising PW11Zn and 2-aminoterephthalic acid in octane–[BIMIM]PF6 biphasic system at 50 oC gives 70% DBT conversion in 4 h and 100% in 6 h reaction time; however this has been obtained using a 50-fold molar excess of H2O2 over DBT [26]. In this system, the ionic liquid [BIMIM]PF6 is used to extract the product sulfone from the model diesel. Keggin PMo polyoxometalate immobilized on phosphazene-modified silica PMo/BzPN-SiO2 gives 100% DBT conversion in heptane with 30% H2O2 at H2O2/DBT = 3.0 and 60 oC, but it takes a longer reaction time of 3 h [27]. 5%[PSPy]3PMo12O40/Graphite catalyst (1 wt% in fuel), where PSPy is N-(3-sulfonatepropyl)-pyridinium ion, gives 100% DBT conversion at 50 oC, H2O2/DBT = 3.0 and 1 h reaction time [28], which is close to the performance of our 10.9%HPMo/C catalyst (1 wt% in fuel, H2O2/DBT = 3.0, 60 oC, 0.5 h). However, the 5%[PSPy]3PMo12O40/Graphite catalyst is prepared by a tedious and lengthy hydrothermal procedure [28], whereas our catalyst is easily obtained from off-the-shelf components by a straightforward wet impregnation procedure (Sect. 2.2).

Moreover, the HPMo/C catalyst exhibited excellent recyclability; it was reused eight times in the oxidation of DBT at 60 oC, giving 100% DBT conversion after each use (Fig. 5). After eight catalyst recycles, catalyst productivity amounted to 10 g DBT per 1 g catalyst and 94 g DBT per 1 g HPMo. The successful catalyst reuse indicates no leaching of HPMo from the catalyst. The lack of leaching is further supported by ICP-OES analysis, which found no Mo in aqueous extract from the reaction mixture after reaction, thus supporting true heterogeneous catalysis in the ODS reaction.

*3.4. Kinetics and mechanism*

Further kinetic and DRIFT spectroscopic studies provided an important insight into the mechanism of the ODS reaction catalyzed by carbon-supported HPA.

As expected, the rate of ODS reaction increased with increasing the temperature. For the oxidation of DBT with HPMo/C catalyst, the apparent activation energy was found to be 49 kJ mol-1 in the temperature range 40 – 70 oC (the Arrhenius plot is shown in Fig. 6). This value is high enough to indicate the absence of diffusion limitations in the reaction system, which is also supported by the independence of reaction rate from the stirring speed (Fig. S4).

The oxidation of BT, DBT and DMDBT in the presence of HPMo/C catalyst was found to be first order in benzothiophene (Fig. 4), with the rate constants 0.0223±0.0008, 0.155±0.004 and 0.0981±0.0031 min-1, respectively, at 60 oC and other conditions specified in Fig. 4. The reaction was also found to be first order in the catalyst HPMo/C (Fig. 7). Therefore, the ODS reaction in the presence of HPMo/C obeyed the rate equation (1), where *k*ODS is the second-order rate constant. The first order in benzothiophenes has also been reported for the ODS in the presence of [PSPy]3PMo12O40/Graphite [28].

(1)

The first order reaction kinetics up to over 90% DBT conversion (Fig. 4) together with eight successive catalyst recycles at 100% DBT conversion each (Fig. 5) evidence against any significant catalyst deactivation during the ODS reaction.

It was found that a 4-fold decrease in concentration of H2O2 in the aqueous phase of the heptane-H2O system did not affect the rate of DBT oxidation, which indicates zero reaction order in H2O2. This was monitored by changing the heptane/H2O ratio from 98/2 to 93/7 v/v by adding 0.2−0.6 mL of H2O to 10 mL of heptane at a constant amount of H2O2 (1.5 mmol) (Table S1).

Fig. 2 shows the DRIFT spectrum for the spent HPMo/C catalyst (spectrum 3) after eight successive runs presented in Fig. 5. Comparison with the spectrum of the fresh catalyst discussed above (spectrum 2) clearly shows that the spent catalyst retained the HPMo Keggin structure. Also the spent catalyst has some DBT sulfone adsorbed on its surface, as evident from comparison with the reference spectrum of the sulfone adsorbed on activated carbon (spectrum 4). This sulfone, however, did not affect catalyst activity as can be seen from the excellent catalyst reuse (Fig. 5).

Another important piece of evidence regarding the active state of HPA/C catalysts in the ODS system was obtained from DRIFT spectra of catalyst samples pretreated with H2O2 (Fig. 8 and Fig. S11). Generally, 31P NMR is a very effective tool for such analysis [12,14,15]. But carbon-supported HPAs give broad 31P MAS NMR peaks [33]; DRIFTS can provide more detailed information in this case. The catalysts were treated with 30% H2O2 under conditions similar to those used for the ODS reaction, except in the absence of benzothiophenes. HPMo/C and HPW/C samples (0.1 g) were stirred with 30% H2O2 (1.5 mmol) in heptane (5 mL) at room temperature or 60 oC for 10 min, filtered off using a Buchner funnel and dried under vacuum to afford a dry, black powder. The samples were ground with KBr, and their DRIFT spectra were measured against a mixture of KBr and activated carbon as the background. As seen in Fig. 8, the HPMo/C catalyst treated with H2O2 at room temperature as well as at 60 oC has exactly the same infrared spectrum as the fresh HPMo/C catalyst in the range of 700-1100 cm-1, characteristic of the Keggin structure. This clearly shows that the structure remains intact after the H2O2 treatment. The same result was obtained for HPW/C catalyst (Fig. S11).

Overall, our DRIFTS data provide an important insight into the mechanism of ODS reaction over carbon-supported HPA catalysts, specifically regarding the active peroxo species involved in this reaction. It is evident that there is a clear difference between homogeneous and heterogeneous ODS based on the H2O2/POM redox system. In homogeneous ODS and alkene epoxidation systems (including truly homogeneous as well as biphasic systems), it has been demonstrated that interaction between the Keggin POM and H2O2 leads to Keggin structure breakdown to form peroxo complexes (observed by 31P and 183W NMR), which act as the active oxidizing species in these reactions ([12,14,15] and references therein). Conversely, in the carbon-supported HPA catalysts, the Keggin structure is preserved due to strong interaction between molecularly dispersed HPA and the carbon surface. In this case, the active oxidizing species could be transient peroxo complexes formed between the intact Keggin polyanions and H2O2.

The vast chemistry of peroxo complexes of Mo(VI) and W(VI), including general patterns of their structure and reactivity, has been thoroughly reviewed by Dickman and Pope [46]. It is recognized that the η2-O22- peroxo ligand is a π-donor like the oxo ligand [46], and the terminal metal-oxo groups M=O can be converted into η2-peroxo groups M(O2)2-. Thus the terminal bonds WVI=O converted into η2-peroxo groups WVI(O2)2- have been found in the lacunary Keggin polyanion [(CoIIO4)W11O31(O2)4]10- [46]. It may therefore be suggested that the terminal Mo=O and W=O bonds in the Keggin HPMo and HPW can interact reversibly with H2O2 to form the corresponding labile η2-peroxo complexes without breaking the Keggin structure on the carbon surface, as shown in Scheme 2. These complexes can then act as the active oxidizing species in ODS reaction. Strong HPA adsorption on the carbon support is thought to be the key to high stability of the Keggin unit in the presence of H2O2. It may not be the case with less adsorbing supports and weaker POM frameworks. Thus, the lacunary [PW11O39]7- unit within a solid hybrid catalyst has been reported to decompose upon interaction with H2O2 to form peroxo WVI species [47].



**Scheme 2.** Formation of peroxo moiety in metal-oxygen octahedron within Keggin HPA supported on activated carbon (M = MoVI or WVI).

The proposed reaction scheme for the oxidation of DBT in model diesel fuel by H2O2 catalyzed by HPA/C is shown in Scheme 3. In the initial step (**1**), the HPA reacts with H2O2 to form the peroxo species on the carbon surface. This species then oxidizes DBT to form DBT sulfone and the initial oxo species via step (**2**), thus completing the catalytic cycle. DBT sulfoxide may be formed in step (**2**), but probably does not desorb, being readily oxidized to DBT sulfone on the catalyst surface. In heptane-H2O system, it was not possible to monitor the sulfoxide and sulfone during reaction because neither was soluble; however intermediate formation of DBT sulfoxide has been observed in different solvents [48]. Sulfur-free diesel fuel can be separated from the solid catalyst and precipitated sulfone by filtration. The sulfone can be separated from the catalyst by solvent extraction (e.g., with toluene), and the catalyst can be reused.



**Scheme 3.** Reaction scheme for oxidation of DBT by H2O2 catalyzed by HPA/C (M = MoVI or WVI).



**Scheme 4.** Proposed step-by-step mechanism for DBT oxidation catalyzed by HPMo/C.

Reflecting on the kinetic and DRIFTS results, the mechanism of oxidation of DBT catalyzed by HPMo/C may be represented by Scheme 4. In step (1), P(Mo=O)12, the initial HPMo/C catalyst, reacts with hydrogen peroxide to form the peroxo species {P(Mo=O)12-m[Mo(O22-)m]} as shown above in Scheme 2. From the DRIFTS results, the Keggin structure remains intact in the presence of H2O2, which indicates that this peroxo species is labile and step (1) is probably equilibrated. This is followed by the oxidation of DBT to sulfoxide (step 2) and further to sulfone (step 3) by the peroxo species. As only traces of sulfoxide were found in the reaction products, it can be suggested that the oxidation of DBT to sulfoxide is the rate limiting step followed by fast oxidation of the sulfoxide adsorbed on the catalyst surface to sulfone. It has been reported that the sulfoxide is oxidised faster that DBT [49]. Finally, the active peroxo species is rapidly regenerated in step (4). Fast formation of the peroxo species in step (1) and its fast regeneration in step (4) will keep the concentration of the peroxo species constant as long as the amount of H2O2 in the reaction mixture exceeds the amount of HPMo. Consequently, as long as this stands, the ODS reaction will be first order in DBT and zero order in H2O2. This is indeed the case as can be seen from the data in Fig. 4 and Table S1.

The order of HPA catalytic activity in the ODS reaction, HPMo > HPW > HSiW, is directly correlated with the rate of oxygen (17O) exchange of Keggin heteropoly anions in aqueous solution, which increases in the order of addenda atoms W(VI) < Mo(VI), as determined from 17O NMR [50]. This is in agreement with the proposed mechanism via steps (1) – (4) because all these steps involve oxygen atom transfer to and from the addenda atoms, which can be expected to be more feasible for Mo(VI) than W(VI).

**4. Conclusions**

This work demonstrates that Keggin-type heteropoly acids supported on activated carbon are active catalysts for oxidative desulfurization (ODS) of model diesel fuel in a biphasic heptane-H2O system using 30% H2O2 as an oxidant. The carbon-supported HPA catalysts are easily obtained from off-the-shelf components by a straightforward wet impregnation procedure. The catalytic activity of HPA/C decreases in the order of HPA: H3PMo12O40 > H3PW12O40 > H4SiW12O40. The most active catalyst, H3PMo12O40/C, removes 100% of benzothiophenes from model diesel fuel at 60 oC and could be recovered and reused many times without loss of activity. This catalyst outperforms other recently reported heterogeneous catalysts for ODS in similar systems. Kinetic and DRIFTS studies suggest a mechanism for heterogeneous ODS reaction on carbon-supported HPAs, which is different from that for a homogeneously-catalyzed ODS reaction. Carbon-supported HPAs, in contrast to their homogeneous counterparts, retain their Keggin structure and form labile active peroxo species without destruction of the HPA Keggin unit, which oxidize benzothiophenes to the corresponding sulfoxides and sulfones.

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**References**

1. I. V. Babich, J. A. Moulijn, Science and technology of novel processes for deep desulfurization of oil refinery streams, Fuel 82 (2003) 607-631.
2. R. Prins, Hydrotreating, in Handbook of Heterogeneous Catalysis, ed. G. Ertl, H. Knözinger, F. Schüth, J. Weitkamp, Vol. 6, Wiley-VCH, 2008, p. 2695-2718.
3. A.W. Bhutto, R. Abro, S. Gao, T. Abbas, X. Chen, G. Yu, Oxidative desulfurization of fuel oils using ionic liquids: A review, J. Taiwan Inst. Chem. Eng. 62 (2016) 84–97.
4. Z. Jiang, H. Lü, Y. Zhang and C. Li, Oxidative desulfurization of fuel oils, Chin. J. Catal. 32 (2011) 707-715.
5. U. Domańska, K. Walczak, M. Królikowski, Extraction desulfurization process of fuels with ionic liquids, J. Chem. Thermodyn. 77 (2014) 40–45.
6. J. Bu, G. Loh, C.G. Gwie, S. Dewiyanti, M. Tasrif, A. Borgna, Desulfurization of diesel fuels by selective adsorption on activated carbons: Competitive adsorption of polycyclic aromatic sulfur heterocycles and polycyclic aromatic hydrocarbons, Chem. Eng. J. 166 (2011) 207–217.
7. D. Boniek, D. Figueiredo, A.F.B. dos Santos, M.A. de Resende Stoianoff, Biodesulfurization: A mini review about the immediate search for the future technology, Clean Technol. Environ. Policy 17 (2014) 29–37.
8. F.M. Collins, A.R. Lucy, C. Sharp, Oxidative desulphurisation of oils via hydrogen peroxide and heteropolyanion catalysis, J. Mol. Catal. A Chem. 117 (1997) 397–403.
9. C. Komintarachat, W. Trakarnpruk, Oxidative desulfurization using polyoxometalates, Ind. Eng. Chem. Res. 45 (2006) 1853–1856.
10. A.F. Shojaei, M.A. Rezvani, M.H. Loghmani, Comparative study on oxidation desulphurization of actual gas oil and model sulfur compounds with hydrogen peroxide promoted by formic acid: Synthesis and characterization of vanadium containing polyoxometalate supported on anatase crushed nanoleaf, Fuel Process. Technol. 118 (2014) 1–6.
11. A.E.S. Choi, S. Roces, N. Dugos, M.W. Wan, Oxidation by H2O2 of bezothiophene and dibenzothiophene over different polyoxometalate catalysts in the frame of ultrasound and mixing assisted oxidative desulfurization, Fuel 180 (2016) 127–136.
12. M. Craven, R. Yahya, E.F. Kozhevnikova, C.M. Robertson, A. Steiner, I.V. Kozhevnikov, Alkylaminophosphazenes as efficient and tuneable phase-transfer agents for polyoxometalate-catalysed biphasic oxidation with hydrogen peroxide, ChemCatChem. 8 (2016) 200–208.
13. I.V. Kozhevnikov, Catalysts for fine chemical synthesis: Catalysis by polyoxometalates, Wiley, West Sussex, 2002.
14. L. Salles, C. Aubry, R. Thouvenot, F. Robert, C. Doremieux-Morin, G. Chottard, H. Ledon, Y. Jeannin, J.-M. Bregeault, 31P and 183W NMR spectroscopic evidence for novel peroxo species in the H3[PW12O40].yH2O/H2O2 system. Synthesis and X-ray structure of tetrabutylammonium (µ-hydrogen phosphato)bis(µ-peroxo)bis(oxoperoxotungstate)(2-): A catalyst of olefin epoxidation in a biphase medium, Inorg. Chem. 33 (1994) 871–878.
15. D.C. Duncan, R.C. Chambers, E. Hecht, C.L. Hill, Mechanism and dynamics in the H3[PW12O40]-catalyzed selective epoxidation of terminal olefins by H2O2. Formation, reactivity, and stability of {PO4[WO(O2)2]4}3-, J. Am. Chem. Soc. 117 (1995) 681–691.
16. C. Venturello, J.C.J. Bart, M. Ricci, A new peroxotungsten heteropoly anion with special oxidizing properties: Synthesis and structure of tetrahexylammonium tetra(diperoxotungsto)phosphate(3-), J. Mol. Catal. 32 (1985) 107–110.
17. C. Venturello, M. Gambaro, A convenient catalytic method for the dihydroxylation of alkenes by hydrogen peroxide, Synthesis 4 (1989) 295–297.
18. C. Venturello, R. D’Aloisio, Quaternary ammonium tetrakis(diperoxotungsto)phosphates(3-) as a new class of catalysts for efficient alkene epoxidation with hydrogen peroxide, J. Org. Chem. 53 (1988) 1553–1557.
19. Y. Ishii, K. Yamawaki, T. Ura, H. Yamada, T. Yoshida, M. Ogawa, Hydrogen peroxide oxidation catalyzed by heteropoly acids combined with cetylpyridinium chloride: Epoxidation of olefins and allylic alcohols, ketonization of alcohols and diols, and oxidative cleavage of 1,2-diols and olefins, J. Org. Chem. 53 (1988) 3587–3593.
20. R. Yahya, M. Craven, E.F. Kozhevnikova, A. Steiner, P. Samunual, I. V. Kozhevnikov, D.E. Bergbreiter, Polyisobutylene oligomer-bound polyoxometalates as efficient and recyclable catalysts for biphasic oxidations with hydrogen peroxide, Catal. Sci. Technol. 5 (2015) 818–821.
21. M. Craven, R. Yahya, E.F. Kozhevnikova, R. Boomishankar, C.M. Robertson, A. Steiner, I.V. Kozhevnikov, Novel polyoxometalate-phosphazene aggregates and their use as catalysts for biphasic oxidations with hydrogen peroxide, Chem. Commun. 49 (2013) 349–351.
22. M. Zhang, W. Zhu, H. Li, M. Li, S. Yin, Y. Li, Y. Wei, H. Li, Facile fabrication of molybdenum-containing ordered mesoporous silica induced deep desulfurization in fuel, Colloids Surface A. 504 (2016) 174–181.
23. J. Zhang, A. Wang, Y. Wang, H. Wang, J. Gui, Heterogeneous oxidative desulfurization of diesel oil by hydrogen peroxide: Catalysis of an amphipathic hybrid material supported on SiO2, Chem. Eng. J. 245 (2014) 65–70.
24. J.M. Fraile, C. Gil, J.A. Mayoral, B. Muel, L. Roldán, E. Vispe, S. Calderón, F. Puente, Heterogeneous titanium catalysts for oxidation of dibenzothiophene in hydrocarbon solutions with hydrogen peroxide: On the road to oxidative desulfurization, Appl. Catal. B Environ. 180 (2016) 680–686.
25. M. Vigolo, S. Borsacchi, A. Soraru, M. Geppi, B.M. Smarsly, P. Dolcet, S. Rizzato, M. Carraro, S. Gross, Engineering of oxoclusters-reinforced polymeric materials with application as heterogeneous oxydesulfurization catalysts, Appl. Catal. B Environ. 182 (2016) 636–644.
26. D. Julião, A.C. Gomes, M. Pillinger, R. Valença, J.C. Ribeiro, B. de Castro, I.S. Gonçalves, L. Cunha Silva, S.S. Balula, Zinc-substituted polyoxotungstate@amino-MIL-101(Al) – An efficient catalyst for the sustainable desulfurization of model and real diesels, Eur. J. Inorg. Chem. 2016 (2016) 5114–5122.
27. M. Craven, D. Xiao, C. Kunstmann-Olsen, E. F. Kozhevnikova, F. Blanc, A. Steiner, I. V. Kozhevnikov, Oxidative desulfurization of diesel fuel catalyzed by polyoxometalate immobilized on phosphazene-functionalized silica, Appl. Catal. B: Environmental 2018, 231, 82-91.
28. W. Jiang, D. Zheng, S. Xun, Y. Qin, Q. Lu, W. Zhu, H. Li, Polyoxometalate-based ionic liquid supported on graphite carbon induced solvent-free ultra-deep oxidative desulfurization of model fuels, Fuel 190 (2017) 1–9.
29. Y. Izumi, K. Urabe, Catalysis of heteropoly acids entrapped in activated carbon, Chem. Lett*.* (1981) 663-666.
30. Y. Izumi, K. Urabe, M. Onaka, Zeolite, clay and heteropoly acid in organic reactions, Kodansha/VCH, Tokyo, 1992; p. 99-161.
31. S.M. Kulikov, M.N. Timofeeva, I.V. Kozhevnikov, V.I. Zaikovskii, L.M. Plyasova, I.A. Ovsyannikova, Adsorption of heteropoly acid H4SiW12O40 from solutions onto porous supports, Izv. Akad. Nauk SSSR, Ser. Khim*.* (1989) 763-768.
32. M.A. Schwegler, H. van Bekkum, N.A. de Munck, Heteropoly acids as catalysts for the production of phthalate diesters, Appl. Catal*.* 74 (1991) 191-204.
33. I.V. Kozhevnikov, A. Sinnema, R.J.J. Jansen, H. van Bekkum, 17O NMR determination of proton sites in heteropoly acid H3PW12O40. 31P, 29Si and 17O NMR, FTIR and XRD study of H3PW12O40 and H4SiW12O40 supported on carbon, Catal. Lett*.* 27 (1994) 187-197.
34. F. Lefebvre, P. Dupont, A. Auroux, Study of the acidity of H3PWI2040 supported on activated carbon by microcalorimetry and methanol dehydration reaction, React. Kinet. Catal. Lett. 55 (1995) 3-9.
35. Study of the acidity of carbon supported and unsupported heteropolyacid catalysts by ammonia sorption microcalorimetry, F.X. Liu-Cai, B. Sahut, E. Faydi, A. Auroux, G. Herve, Appl. Catal. A 185 (1999) 75-83.
36. M.S. Kaba, M.A. Barteau, W.Y. Lee, I.K. Song, Nanoscale characterization of acid properties of heteropolyacids byscanning tunneling microscopy and tunneling spectroscopy, Appl. Catal. A 194 -195 (2000) 129-136.
37. R. Neumann, M. Levin, The selective aerobic oxidative dehydrogenation of alcohols and amines catalysed by a supported molybdenum-vanadium heteropolyanion salt Na5PV2Mo10040, J. Org. Chem*.* 56 (1991) 5707-5710.
38. S. Fujibayashi, K. Nakayama, M. Hamamoto, S. Sakaguchi, Y. Nishiyama, Y. Ishii, An efficient aerobic oxidation of various organic compounds catalyzed by mixed addenda heteropolyoxometalates containing molybdenum and vanadium, J. Mol. Catal. A 110 (1996) 105-117.
39. L. Xu, E. Boring, C.L. Hill, Polyoxometalate-modified fabrics: New catalytic materials for low-temperature aerobic oxidation, J. Catal*.* 195 (2000) 394-405.
40. C. Rocchiccioli-Deltcheff, M. Fournier, R. Franck, R. Thouvenot, Evidence for anion-anion interactions in molybdenum(VI) and tungsten(VI) compounds related to the Keggin structure, Inorg. Chem. 22 (1983) 207-216.
41. G. M. Brown, M.-R. Noe-Spirlet, W. R. Busing, H. A. Levy, Dodecatungstophosphoric acid hexahydrate, (H5O2+)3(PW12O403-). The true structure of Keggin's `pentahydrate' from single-crystal X-ray and neutron diffraction data, Acta Cryst*.* B33 (1977) 1038-1046.
42. I.V. Kozhevnikov, Catalysis by heteropoly acids and multicomponent polyoxometalates in liquid-phase reactions, Chem. Rev. 98 (1998) 171–198.
43. G. Yu, S. Lu, H. Chen, Z. Zhu, Diesel fuel desulfurization with hydrogen peroxide promoted by formic acid and catalyzed by activated carbon, Carbon 43 (2005) 2285–2294.
44. X. Jiang, H. Li, W. Zhu, L. He, H. Shu, J. Lu, Deep desulfurization of fuels catalyzed by surfactant-type decatungstates using H2O2 as oxidant, Fuel 88 (2009) 431–436.
45. J. Zhang, A. Wang, X. Li, X. Ma, Oxidative desulfurization of dibenzothiophene and diesel over [Bmim]3PMo12O40, J. Catal. 279 (2011) 269–275.
46. M.H. Dickman, M.T. Pope, Peroxo and superoxo complexes of chromium, molybdenum, and tungsten, Chem. Rev.94 (1994) 569-584.
47. Z.X. Zhang, W. Zhao, B.C. Ma, Y. Ding, Catal. Commun. The epoxidation of olefins catalyzed by a new heterogeneous polyoxometalate-based catalyst with hydrogen peroxide, 12 (2010) 318-322.
48. E. Rafiee, N. Rahpeyma, Selective oxidation of sulfurs and oxidation desulfurization of model oil by 12-tungstophosphoric acid on cobalt-ferrite nanoparticles as magnetically recoverable catalyst. Chin. J. Catal. 36 (2015) 1342-1349.
49. Kinetics and mechanism of the oxidation of dibenzothiophene in hydrocarbon solution. Oxidation by aqueous hydrogen peroxide-acetic acid mixtures, B.N. Heimlich, T.J. Wallace, Terahedron, 22 (I966) 3571-3579.
50. M. A. Eedotov, R. I. Maksimovskaya, D. U. Begalieva, A. K. Il'yasova, 17O NMR spectra and isotope exchange of oxygen in aqueous solutions of phosphorus-vanadium-tungsten heteropoly acids, Russ. Chem. Bull. 29 (1980) 1025-1028.

**Table 1**

Information about HPA/C catalysts.

|  |  |  |  |
| --- | --- | --- | --- |
| Catalyst a | *S*BET bm2g-1 | Pore volume c  cm3g-1 | Pore diameter d  Å |
| Activated carbon | 977 | 0.89 | 35 |
| 10.9%HPMo/C | 1072 | 0.94 | 34 |
| 11.0%HPW/C | 1006 | 0.85 | 33 |
| 12.7%HSiW/C | 1002 | 0.88 | 34 |

a HPA loading calculated from Mo and W content determined by ICP-OES analysis. b BET surface area. c Single point pore volume. d Average pore diameter.

**Table 2**

Oxidation of benzothiophenes by H2O2 in heptane-H2O (98:2 v/v) system using HPA/C catalysts.a

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Entry | HPA/C | Substrate | H2O2/Sb  mol/mol | Time  h | Conversionc  % |
| 1 | 10.9%HPMo/C | DBT | 3.0 | 0.5 | 100d |
| 2 | 11.0%HPW/C | DBT | 3.0 | 3.0 | 89 |
| 3 | 12.7%HSiW/C | DBT | 3.0 | 3.0 | 49 |
| 4 | 10.9%HPMo/C | DBT | 2.0 | 1.5 | 98±2e |
| 5 | 10.9%HPMo/C | DBT | 0f | 3.0 | 5g |
| 6 | Activated carbonh | DBT | 3.0 | 3.0 | 30 |
| 7 | 10.9%HPMo/C | BT | 3.0 | 2.0 | 100i |
| 8 | 10.9%HPMo/C | DMDBT | 3.0 | 1.0 | 100d |

a 60 oC, HPA (0.0041mmol), substrate (BT, DBT or DMDBT, 0.50 mmol), H2O2 (1.50 mmol, 0.166 mL of 30% H2O2 (9.05 M) unless stated otherwise), dodecane (GC standard, 0.40 mmol) and heptane (10 mL); stirring speed 1500 rpm. b H2O2/substrate molar ratio. c Substrate conversion to sulfone. d 80% H2O2 efficiency from post reaction titration with KMnO4; the yield of DBT sulfone 100±13%. e Reaction at stoichiometric molar ratio H2O2/S = 2.0; hydrogen peroxide efficiency 98±2% (average of three runs). f 0.2 mL H2Oadded instead of 30% H2O2. g The percentage of DBT adsorbed on the catalysts. h 0.024 g of activated carbon added instead of HPA/C catalyst. i ≥75% H2O2 efficiency.

**Table 3**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| Catalyst | Model diesel | H2O2/DBT  mol/mol | T  oC | Time  h | DBT conversion  % | Reference |
| HPMo/C | heptane | 3.0 | 60 | 0.5 | 100 | This work |
| Ti(IV)/SiO2 | isooctane | 5.0 | 60 | 8 | 99 | [22] |
| Zr(IV)/PMMAb | octane | 3.5 | 65 | 24 | 84 | [23] |
| PW11Zn-MOFc | octane | 50 | 50 | 6 | 100 | [24] |
| PMo/BzPN-SiO2d | heptane | 3.0 | 60 | 3 | 100 | [25] |
| [PSPy]PMo/Ce | octane | 3.0 | 50 | 1 | 100 | [26] |

Comparison of heterogeneous catalysts for oxidation of DBT by H2O2 in model diesel fuel.a

a Biphasic systems comprising a model diesel phase and aqueous H2O2 phase. b Zr(IV) oxoclusters in poly(methylmethacrylate) matrix (PMMA). c MOF comprising PW11Zn polyoxometalate and 2-aminoterephthalic acid in octane–[BIMIM]PF6 biphasic system. d PMo polyoxometalate immobilized on phosphazene-functionalized silica. e 5%[PSPy]3PMo12O40/Graphite (PSPy = N-(3-sulfonatepropyl)-pyridinium ion), 1% in reaction mixture.

**Figure captions**

**Fig. 1.** Nitrogen adsorption (solid markers) and desorption (open markers) for Darco KB-B activated carbon (triangles) and 10.9%HPMo/C catalyst (circles).

**Fig. 2.** DRIFT spectra of (1) bulk H3PMo12O40, (2) fresh 10.9%HPMo/C catalyst, (3) spent 10.9%HPMo/C catalyst after 8 successive runs of DBT oxidation and (4) DBT sulfone adsorbed on activated carbon (powdered sample mixtures with KBr; (1) versus pure KBr background, (2) – (4) versus a mixed KBr + carbon background).

**Fig. 3.** Effect of HPA on the oxidation of DBT (0.50 mmol) by H2O2 (1.5 mmol) in heptane (10 mL) catalyzed by HPA/C at 60 oC (HPA, 0.0041 mmol).

**Fig. 4.** Plot of substrate conversion (*x*) versus reaction time and first-order plot ln (1 – *x*) = –*kt* for oxidation of BT, DBT and 4,6-DMDBT (0.50 mmol) by H2O2 (1.5 mmol) in heptane (10 mL) catalyzed by 10.9%HPMo/C (0.0685 g, 1.0 wt%, 0.0041 mmol HPMo) at 60 oC.

**Fig. 5.** Catalyst reuse in oxidation of DBT (0.50 mmol, 1.3 wt% in reaction mixture) by H2O2 (1.5 mmol) in heptane (10 mL) catalyzed by 10.9%HPMo/C (0.0685 g, 1.0 wt%, 0.0041 mmol HPMo) at 60 oC, 1 h.

**Fig. 6.** Arrhenius plot for oxidation of DBT (0.50 mmol) by H2O2 (1.5 mmol) catalyzed by 10.9%HPMo/C (0.0343 g, 0.0020 mmol HPMo).

**Fig. 7.** Effect of catalyst amount on the rate of DBT (0.50 mmol) oxidation by H2O2 (1.5 mmol) catalyzed by 10.9%HPMo/C at 60 oC (*k* is the first-order rate constant).

**Fig. 8.** DRIFT spectra of (1) bulk H3PMo12O40, (2) fresh 10.9%HPMo/C catalyst, (3) 10.9%HPMo/C catalyst treated with H2O2 at 60 oC and (4) 10.9%HPMo/C catalyst treated with H2O2 at 20 oC (powdered sample mixtures with KBr; (1) versus pure KBr background, (2) – (4) versus a mixed KBr + carbon background).

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-0.2

-0.1

0.0

0.1

0.2

0.3

0.4

0.5

0.6

0.7

0.8

Absorbance (a. u.)

800

1000

1200

Wavenumbers (cm-1)

1100

900

600

700

1066

792

964

1062

811

960

871

884

**4**

**3**

**2**

**1**

**Fig. 8.** DRIFT spectra of (1) bulk H3PMo12O40, (2) fresh 10.9%HPMo/C catalyst, (3) 10.9%HPMo/C catalyst treated with H2O2 at 60 oC and (4) 10.9%HPMo/C catalyst treated with H2O2 at 20 oC (powdered sample mixtures with KBr; (1) versus pure KBr background, (2) – (4) versus a mixed KBr + carbon background).

1. Corresponding author. *E-mail address:* [kozhev@liverpool.ac.uk](mailto:kozhev@liverpool.ac.uk) (I.V. Kozhevnikov). [↑](#footnote-ref-1)